

RABINDRA MAHAVIDYALAYA



Affiliated to The University of Burdwan
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Estd. -1971



Ref. No.....


Date:.....

2.3.1 – Student centric methods, such as experiential learning, participative learning and problem solving methodologies are used for enhancing learning experiences and teachers use ICT –enabled tools including online resources for effective teaching and learning process

Teachers use ICT –enabled tools including online resources for effective teaching and learning process

List of content:

- Documents for various E-resources and Techniques used for the session 2022-23
- Documents for various E-resources and Techniques used for the session 2021-22
- Documents for various E-resources and Techniques used for the session 2020-21


14.08.2024
Principal
Rabindra Mahavidyalaya
Champadanga, Hooghly

Session 2022-23

Teachers use ICT enabled tools for effective teaching-learning process.


4.8.2024
ATTESTED
Principal
Rajendra Kumar
Chandigarh, Punjab



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Estd.- 1971



Ref. No.....

Date.....

Session 2022-23

2.3.2 Teachers use ICT enabled tools for effective teaching-learning process.

Along with offline classes some Departments also use online mode of teaching. Google Meet, Zoom, Youtube links, Jamboard in Google meet and Tesmoz App have been used. For this purpose Desktop, laptop, LCD projector, printer, pen drive, Mobile Smart phone, microphone, scanner, external hard drive, sound system etc were used by the various Departments. Video, PPT, PDF, WORD documents, excel, e-resources, e-journals, e-books, notes etc. study materials were provided to the students during this academic session. E-Journals and e-books were used for effective teaching and learning. A few numbers of ICT enabled classrooms were developed in some Department and also in the college campus as central facilities.

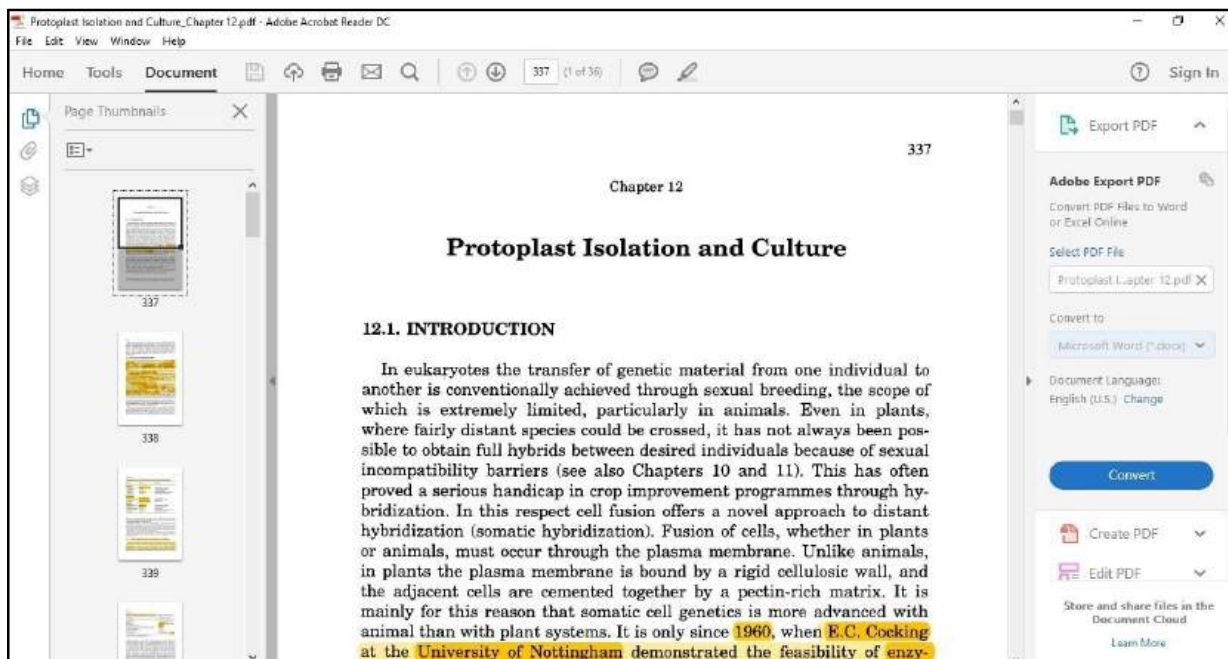
DEPARTMENT OF BOTANY

Snapshots/ screenshots of E-resources and techniques used in 2022-23

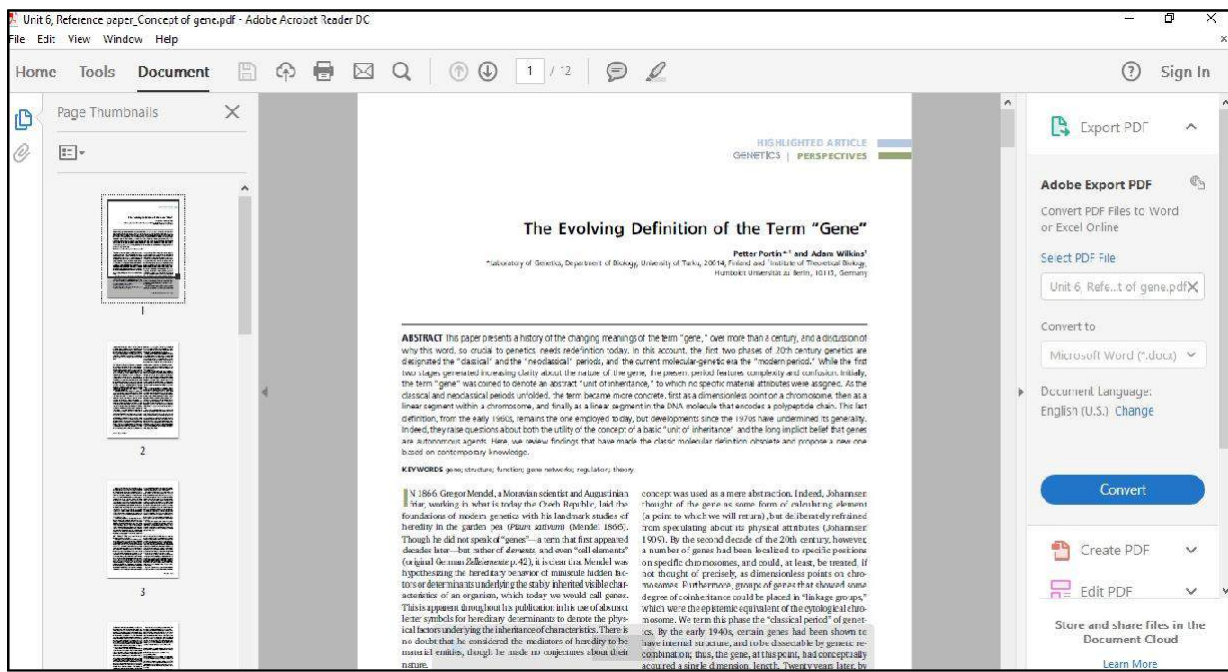


Teachers taking ICT enabled classes taken in classrooms

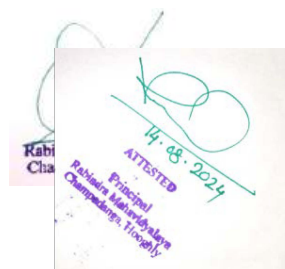
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Screenshot (sample) of E- Book chapter



Screenshot (sample) of E- Journal Resource



LECTURE 3_DIHYBRID CROSS

DIHYBRID TEST CROSS

Round, yellow
 $RrYy$

Wrinkled, green
 $rryy$

Expected proportions for first character
 $Rr \times rr$
Cross
 $\frac{1}{2} Rr$ Round
 $\frac{1}{2} rr$ Wrinkled

Expected proportions for second character
 $Yy \times yy$
Cross
 $\frac{1}{2} Yy$ Yellow
 $\frac{1}{2} yy$ Green

Expected proportions for both characters
 $RrYy \times rryy$

PHENOTYPE & GENOTYPE RATIO:
1 : 1 : 1 : 1

A branch diagram can be used for determining the phenotypes and expected proportions of offspring from a **di-hybrid testcross** ($RrYy \times rryy$).

$\frac{1}{2} Rr$ Round

$\frac{1}{2} rr$ Wrinkled

$\frac{1}{2} Yy$ Yellow

$\frac{1}{2} yy$ Green

$RrYy$ $\frac{1}{2} \times \frac{1}{2} = \frac{1}{4}$ Round, yellow

$Rryy$ $\frac{1}{2} \times \frac{1}{2} = \frac{1}{4}$ Round, green

$rrYy$ $\frac{1}{2} \times \frac{1}{2} = \frac{1}{4}$ Wrinkled, yellow

$rryy$ $\frac{1}{2} \times \frac{1}{2} = \frac{1}{4}$ Wrinkled, green

1634

XRecorder

Screenshot of video (sample)

Department of Defence Studies(2022-23)



Prof. Jana delivered lecture on Submerin Mechanism.



Prof. Adhikari delivered lecture on Granade Mechanism.



Prof. Adhikari delivered lecture on finding true north.

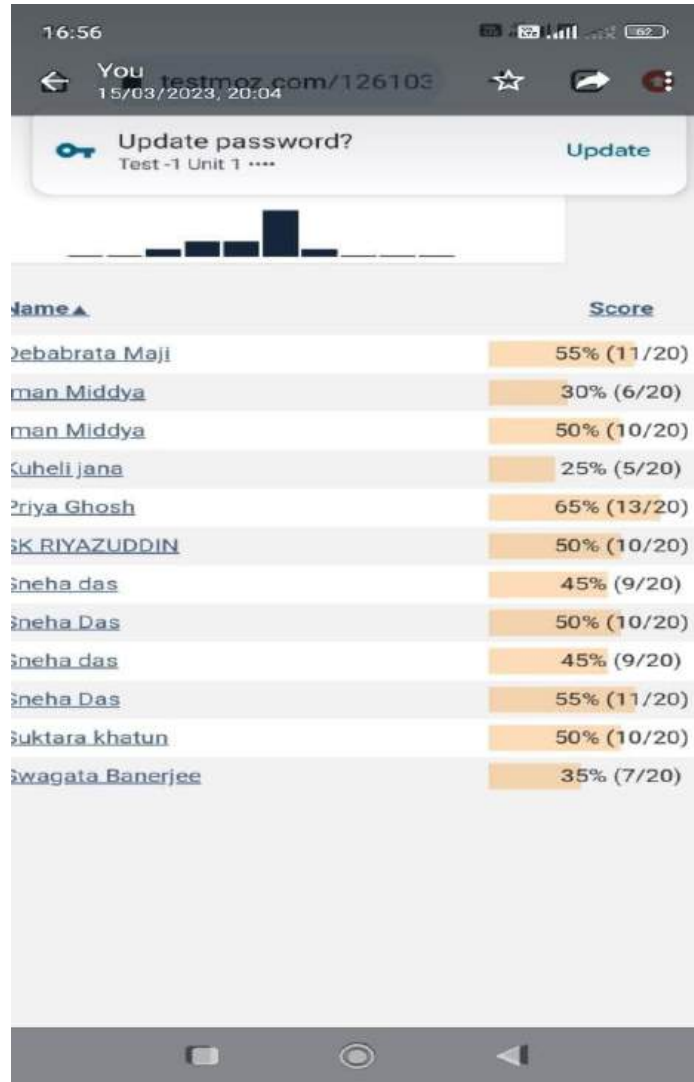


Prof. Jana delivered lecture on Submerin Main parts.

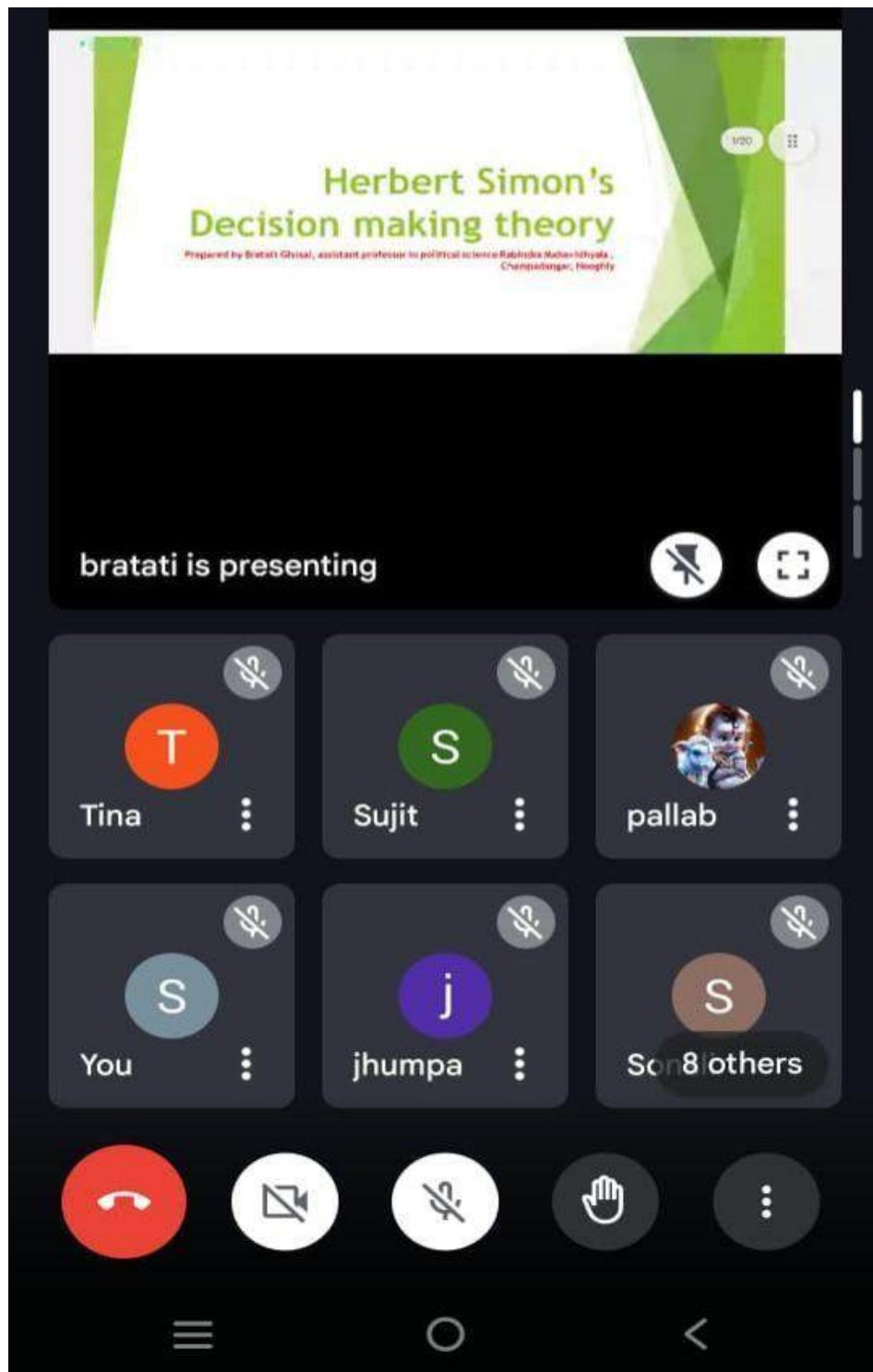


Department of Political Science(2022-2023)

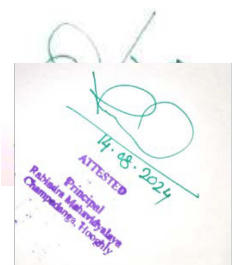
Result of sample class test taken with the use of testmoz app



AT
Pr
Rabindra
Champer
14.08.2024
ATTESTED
Principal
Rabindra Kumar Chatterjee
Champer, Hooghly



Sample ppt presentation for sem -3, paper CC-6



DEPARTMENT OF GEOGRAPHY(2022-2023)



Photo. Ongoing RS GIS class in the Department of Geography –An example of ICT enabled class

Rabind
Champ
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Chandrapur, Hooghly



Figure. Making sense of intricate regional geography mapping using Google Books



Figure. Using Google Books to help understand complex climate diagrams

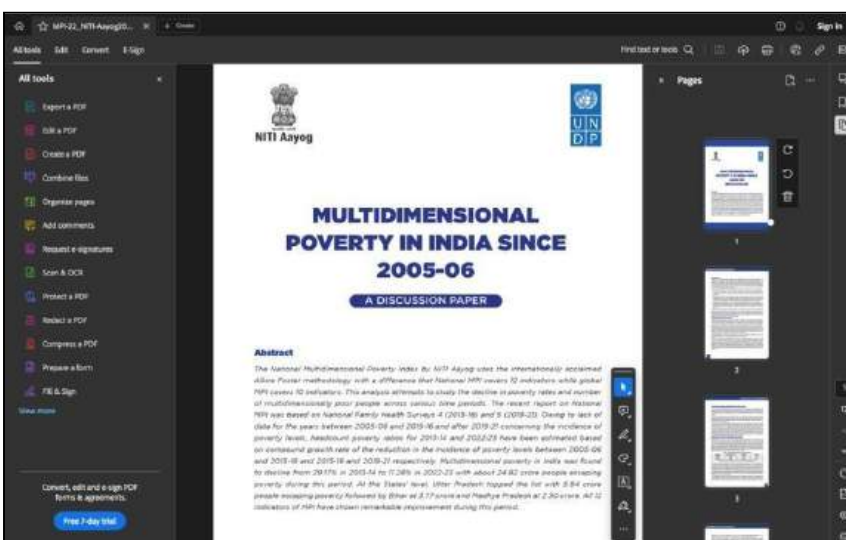
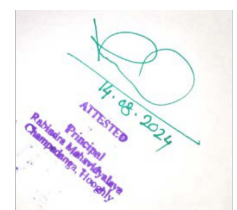
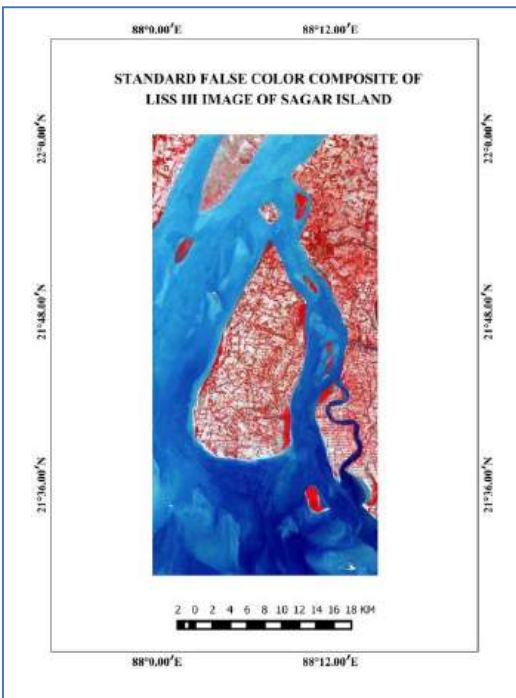
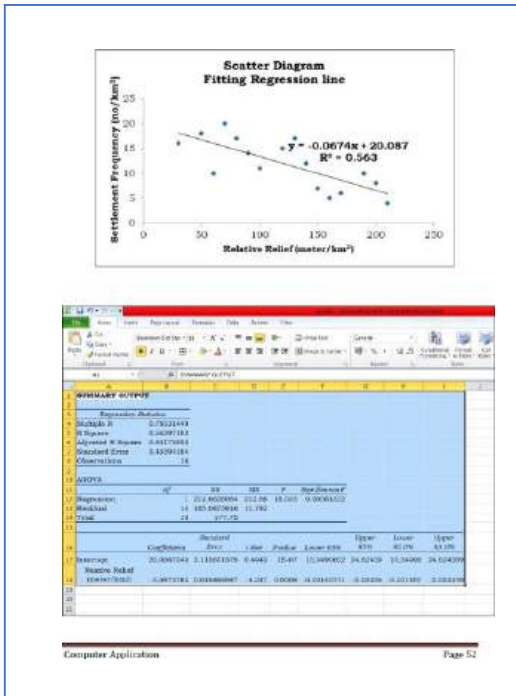


Figure. Sample e-discussion paper for understanding theory of Economic Geography





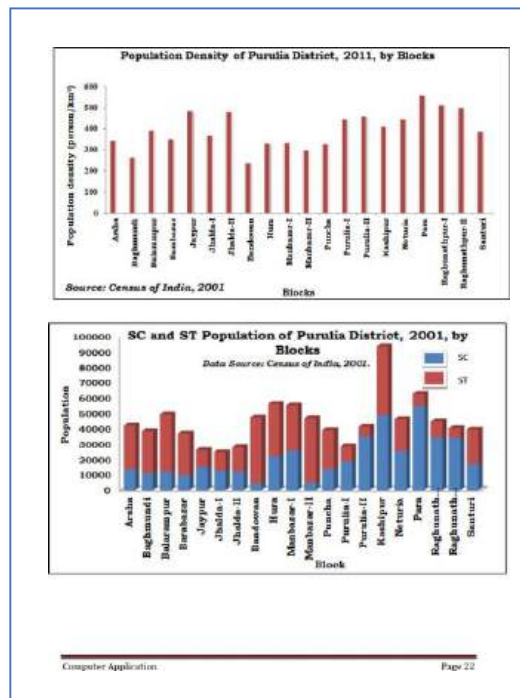
Figures. Students of Geography have done project work using ICT enabled tools

THE UNIVERSITY OF BURDWAN

Quantum GIS

ROLL: NO:
REGN NO: SESSION:

LABORATORY NOTE BOOK
FOR
COMPUTER APPLICATION
REMOTE SENSING AND GIS



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Rabindra Mahapatra
Chandrapur, Hooghly

Department of Chemistry (2022-2023)

Snapshots/ screenshots of E-resources and techniques used

Screenshot (sample) of Resources

PPT Sample

PERIODIC TABLE

By
Dr. Sucheta Joy
Assistant Professor
Department of Chemistry
Rabindra Mahavidyalaya
Champadanga, Hooghly

Electron Configurations

Atomic Radii

The atomic radius is defined as one-half of the distance between covalently bonded atoms.

Effective Nuclear Charge

The effective nuclear charge, Z_{eff} is found by:
$$Z_{\text{eff}} = Z - S$$

where Z is the atomic number and S is a screening constant, usually close to the number of inner electrons.
Note: For simple estimation, we may use the inner number of electrons.

CHEMISTRY HONS. SEM- V
PHYSICAL CHEMISTRY
LESSON 1
BY
DR. DEBASMITA SARDAR
ASSISTANT PROFESSOR
DEPARTMENT OF CHEMISTRY
RABINDRA MAHAVIDYALAYA
CHAMPADANGA, HOOGHLY

Course Code- DSE-1
Course Title- Advanced Physical Chemistry (Theo)
Statistical Thermodynamics

Topics

1. Configuration: Macrostates, microstates and configuration; calculation with harmonic oscillator; variation of W with E ; equilibrium configuration
2. Boltzmann distribution: Thermodynamic probability, entropy and probability; Boltzmann distribution formula (with derivation); Applications to barometric distribution; Partition function, concept of ensemble - canonical ensemble
3. Partition function: molecular partition function and thermodynamic properties; Maxwell's speed distribution; Gibbs' paradox

Introduction

- > Statistical mechanics, one of the pillars of modern physics, describes how macroscopic observations are related to microscopic parameters that fluctuate around an average
- > What are macroscopic properties?
- > Who is father of statistics?
- > What are the meanings of statistics?
- > What is the definition of statistics?
- > What is statistical mechanics used for?
- > What is the aim of statistical thermodynamics?

Statistical thermodynamics
Lectures 7, 8

Quantum \longleftrightarrow Classical

Energy levels \longleftrightarrow Bulk properties

Various forms of energies

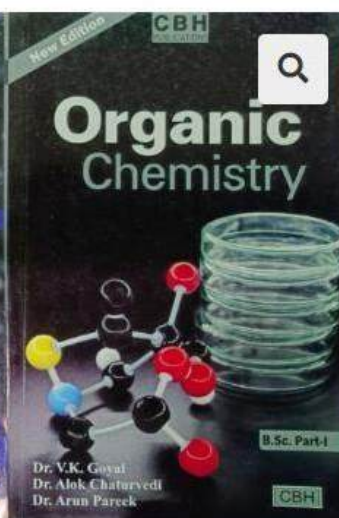
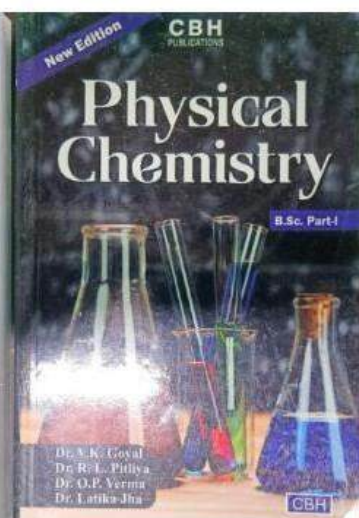
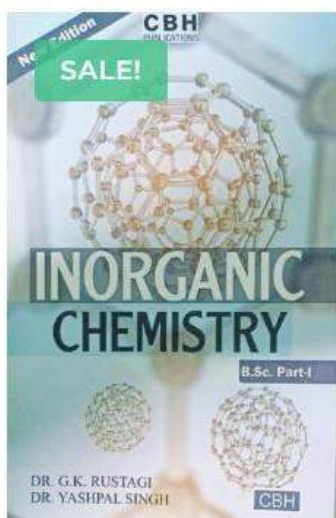
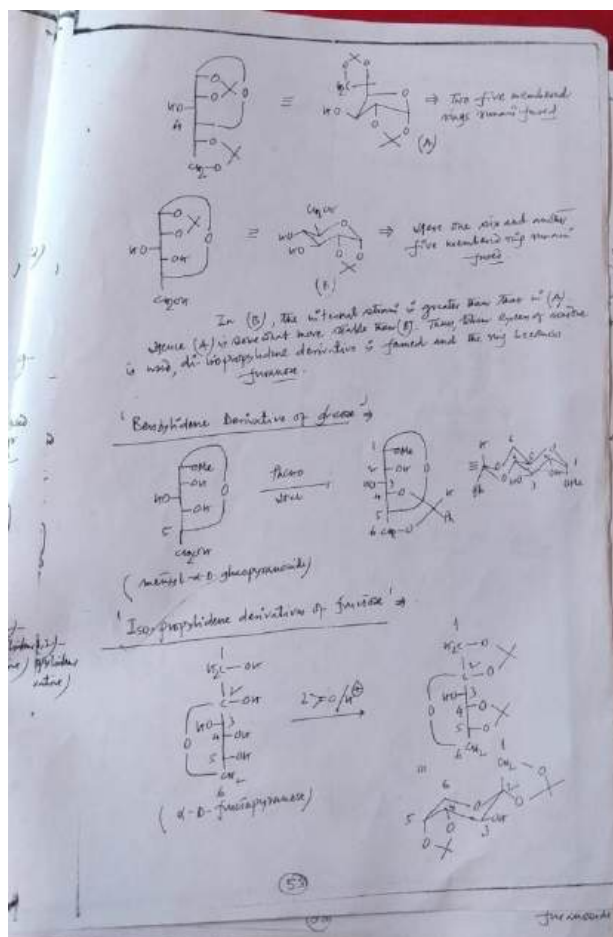
Everything turns out to be controlled by temperature

Why is it useful?

- There are 2 strands of Physical Chemistry:
 - Macroscopic: Thermodynamics and kinetics (involves molecules, atoms, and if so, at what rate?)
 - Microscopic: atomic structure, quantum chemistry and their explanation by quantum mechanics
- These (apparently) unrelated strands can be linked by statistical thermodynamics.
- E.g. we can calculate the enthalpy of a reaction instead of measuring it!

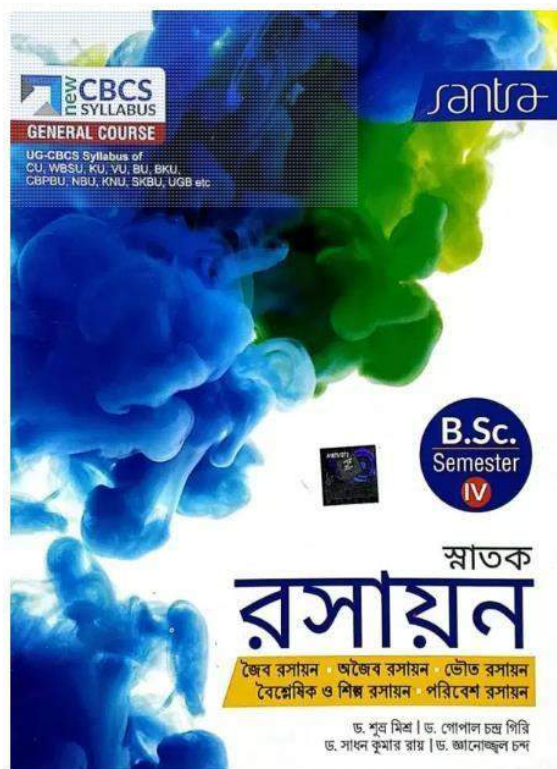
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PDF Sample/ Books



14.08.2024
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Principal
Chandigarh University

Bengali book for general students



ICT Classes



14.08.2024
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Principal
Pabitra Mahapatra
Chandrapur, Hooghly

Class taken by Dr. Sucheta Joy, Assistant Professor, Department of Chemistry



Class taken by Dr. Debasmita Sardar, Assistant Professor, Department of Chemistry



Class taken by Dr. Rabiul Alam, Assistant Professor, Department of Chemistry

14.05.2024
ATTESTED
Principal
Rabiul Alam
Chemistry, Hooghly

Session 2021-22

Teachers use ICT enabled tools for effective teaching-learning process.





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Estd.- 1971




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Date.....

This is to certify that the following ICT enabled tools for effective teaching-learning process was used by the various Departments in the session 2021-2022 at Rabindra Mahavidyalaya, Champadanga, Hooghly, West Bengal


Principal
Rabindra Mahavidyalaya
Champadanga - Hooghly

Dr. Prasanta Bhattacharyya
Principal


14.08.2024
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Principal
Rabindra Mahavidyalaya
Champadanga, Hooghly

Session 2021-22

Teachers use ICT enabled tools for effective teaching-learning process.

Due to Covid-19 entire teaching learning was switched to online mode. Google Meet, Zoom, etc were used for teaching through. desktop, Laptop. Smart Phone etc. Video, PPT, PDF, excel, etc files were shared for study material. E-Journals and e-books were used for effective teaching and learning.

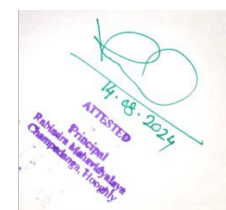
Some screenshots of the same are attached here.

Department of Botany (2021-2022)

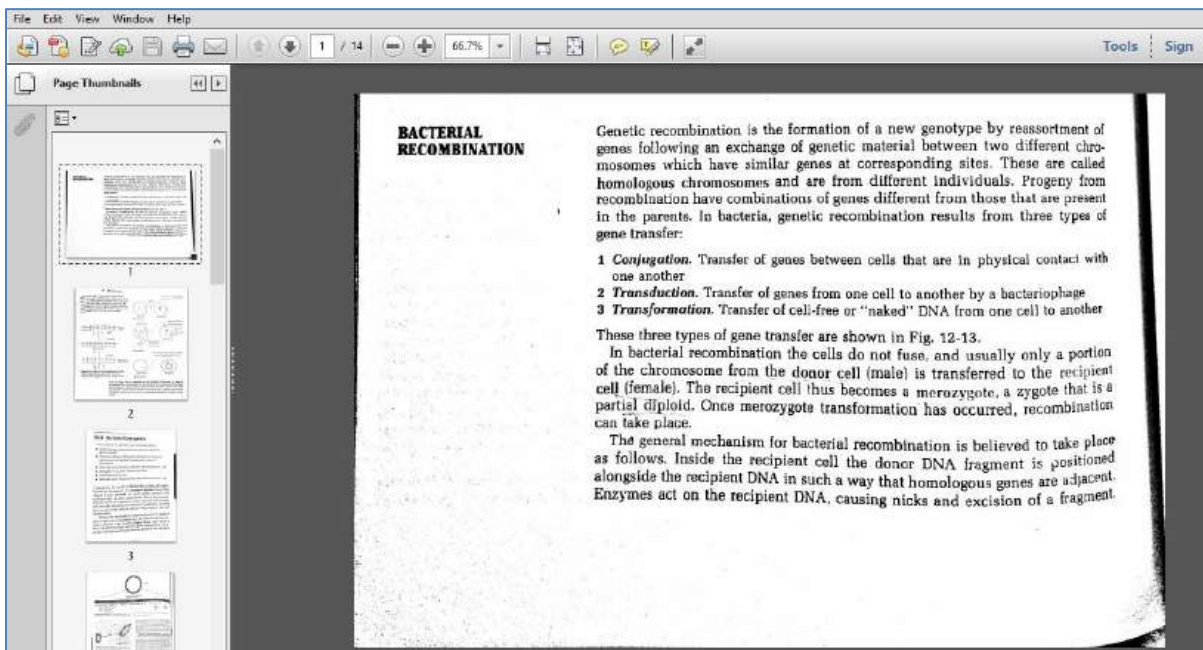
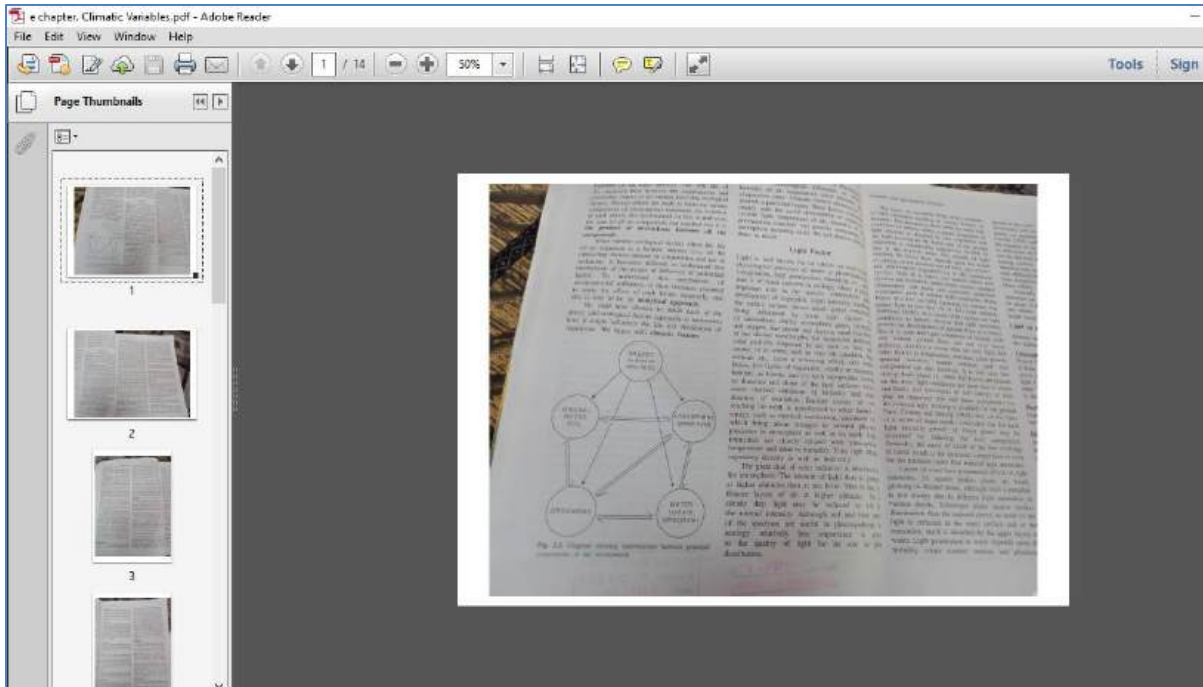
Snapshots/ screenshots of E-resources and techniques used

Screenshot (sample) of E- Journal Resource

The screenshot displays a web browser window showing a Taylor & Francis online journal article. The URL in the address bar is <https://doi.org/10.1080/15592324.2019.1596719?needAccess=true&role=button>. The page layout includes a left sidebar with a 'Details' section showing the journal cover for 'PLANT signaling & behavior', Volume 14, Issue 6, 2019. The main content area features the article title 'Plant defense against virus diseases; growth hormones in highlights' by Waqar Islam, Hassan Naveed, Madiha Zaynab, Zhiqun Huang, and Han Y. H. Chen. The article is categorized as a 'Review'. The abstract states: 'Phytohormones are critical in various aspects of plant biology such as growth regulations and defense strategies against pathogens. Plant-virus interactions retard plant growth through rapid alterations in phytohormones and their signaling pathways. Recent research findings show evidence of how viruses impact upon modulation of various phytohormones affecting plant growth regulations. The opinion is getting stronger that virus-mediated phytohormone disruption and alteration weakens plant defense strategies through enhanced replication and systemic spread of viral particles. These hormones regulate plant-virus interactions in various ways that may involve antagonism and cross talk to modulate small RNA (siRNA) systems. The article aims to highlight the recent research findings elaborating the impact of viruses upon manipulation of phytohormones and virus biology.' The article history shows it was received on 6 March 2019, revised on 20 May 2019, and accepted on 12 March 2019. The keywords listed are: plant defense pathways, ethylene, salicylic acid, jasmonic acid, gibberellins, auxin, cytokinins, decarboxylase, and, isochlorogenic acid. The introduction section begins with: 'Plant viruses utilize numerous strategies that are more conducive for replication and viral spread inside the plant's cellular environment.' The article is published by Taylor & Francis Group.



PDF Sample



14.08.2024

ATTESTED

Principal

Principals

Principals

Principals

Unit 1_1.8_Complex permanent tissue (Xylem).pdf - Adobe Reader

File Edit View Window Help

Page Thumbnails

1

2

3

Core Course (CC- IC): Plant Anatomy and Embryology
Unit 1: Meristematic and permanent tissues
1.8. Complex permanent tissue
Topic: XYLEM

II. জটিল স্থায়ী কলা (Complex Permanent Tissues) : যাকার ও গঠন নিয়ে হওয়া সত্ত্বেও একই বস্তু সম্পর্কযুক্ত এবং একই উদ্দেশ্য থেকে সৃষ্টি হয়েছে এরকম রূপে সংজ্ঞায়িত জটিল স্থায়ী কলা বলে।
এই কলা উদ্ভিদের সংরক্ষণ কলা বা পরিবহন কলা নামে পরিচিত। এই জটিল কলা দুই রকমের, যথা—জাইলেম কলা ও ফ্লোয়েম কলা। এই জাইলেম ও ফ্লোয়েমের মাধ্যমে যথাক্রমে কলা ও খাদ্য বিভিন্ন অংশে পরিবহিত হয়, তাই এদের সংরক্ষণ কলা (conducting tissue) বলে। প্রধানত জাইলেম ও ফ্লোয়েম একত্রে নালিকা বাঁধিল (vascular bundle) গঠন করে।

1) জাইলেম (Xylem) :

● সংজ্ঞা (Definition) : উদ্ভিদের যে জটিল স্থায়ী কলার মাধ্যমে মূল থেকে শোষিত জল ও জলে দ্রবীভূত বিভিন্ন লবণ উদ্ভিদের শাখায় পরিবহিত হয়, তাকে জাইলেম কলা বলে।
এটি গ্রিক শব্দ 'xylem' (- বাই) থেকে এসেছে।

2) গঠন (Structure) :

(i) জাইলেম হল অন্যতম জটিল স্থায়ী কলা।
(ii) এই জাইলেম ও ফ্লোয়েম একত্রে নালিকা বাঁধিল (vascular bundle) গঠন করে অর্থাৎ জাইলেম নালিকা বাঁধিলের একটি অংশ গঠন করে।
(iii) জাইলেম বিভিন্ন প্রকারের মৃত ও সজীব কোষ দ্বারা গঠিত। জাইলেম প্যারেনকাইমা জাইলেমের একত্রে সজীব কলা।
(iv) সাধারণত অংশগ্রহণকারী ট্রাকিয়ার ট্যাপন (tracheary elements), যেনে—ট্রাকিড ও ট্রাকিয়া, জাইলেম প্যারেনকাইমা (xylem parenchyma) এবং জাইলেম তন্তু (xylem fibres) নিয়ে জাইলেম কলা গঠিত।

3) অবস্থান (Occurrence) : একবীজপত্রী ও দ্বিবীজপত্রী উদ্ভিদের মূল, কাণ্ড ও শাখায় নালিকা বাঁধিলে (vascular bundles) জাইলেম কলা অবস্থান করে।

4) উপাদান (Components) : জাইলেম চার প্রকার সজীব ও মৃত কোষ নিয়ে গঠিত, যথা— (A) ট্রাকিড (tracheid), (B) ট্রাকিয়া (trachea), (C) জাইলেম প্যারেনকাইমা (xylem parenchyma) এবং (D) জাইলেম তন্তু (xylem fibres)।
উপরে উপাদানগুলির মধ্যে ট্রাকিড (tracheid) এবং জাইলেম বারিধা (vessel members) অর্থাৎ ট্রাকিয়া (trachea) জল ও খনিজ লবণ পরিবহনে সাহায্য করে, সেই কারণে এদের একত্রে ট্রাকিয়ার উপাদান (tracheary elements) বলে।

(A) ট্রাকিড (Tracheid) :

● সংজ্ঞা (Definition) : পুরু কোষপ্রান্তরযুক্ত, লম্বা, ইঁটালো প্রান্তবিন্দিত মৃত জাইলেম কোষকে ট্রাকিড (tracheid) বলে।

14.08.2024

ATTESTED

Principal

Patana Mahavidyalaya

Chandrapur, Tripura

PPT. Sample

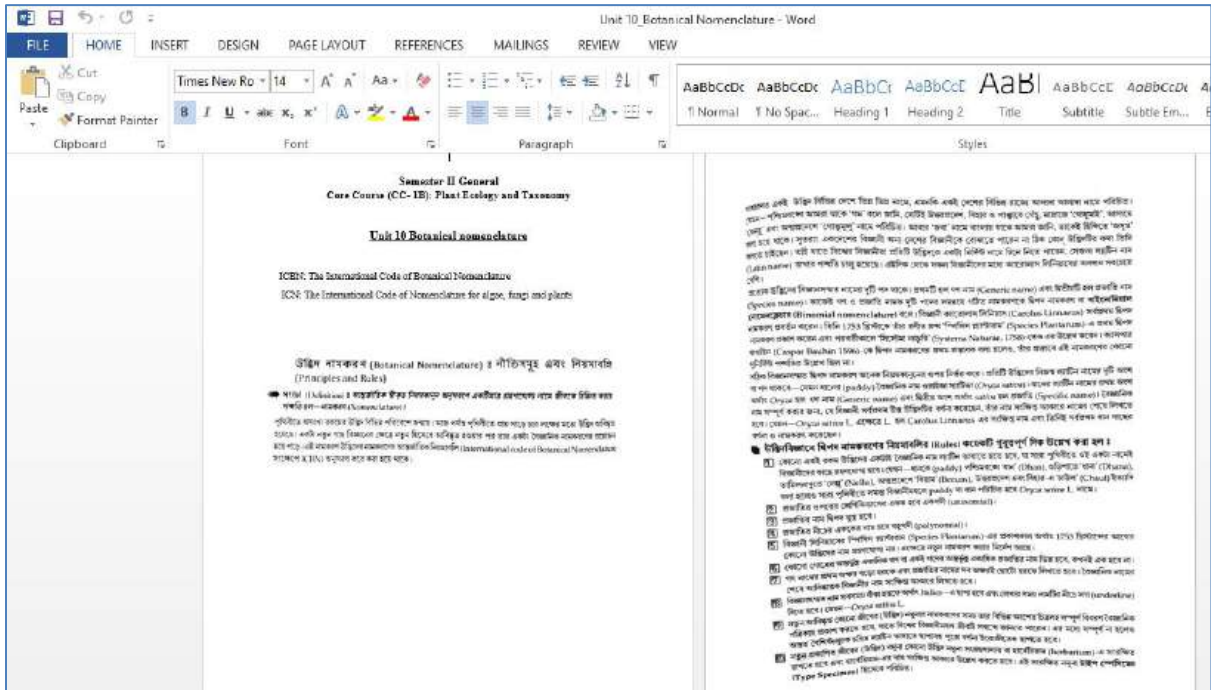
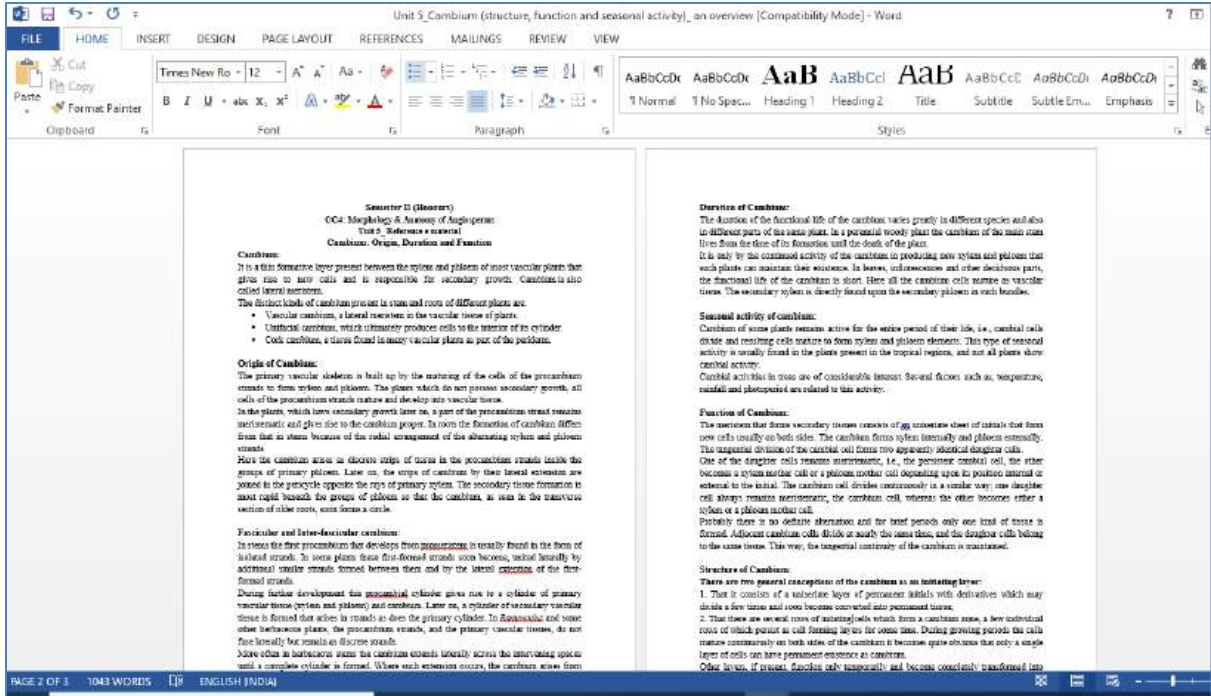
[illegible][illegible]

14.08.2024

ATTESTED

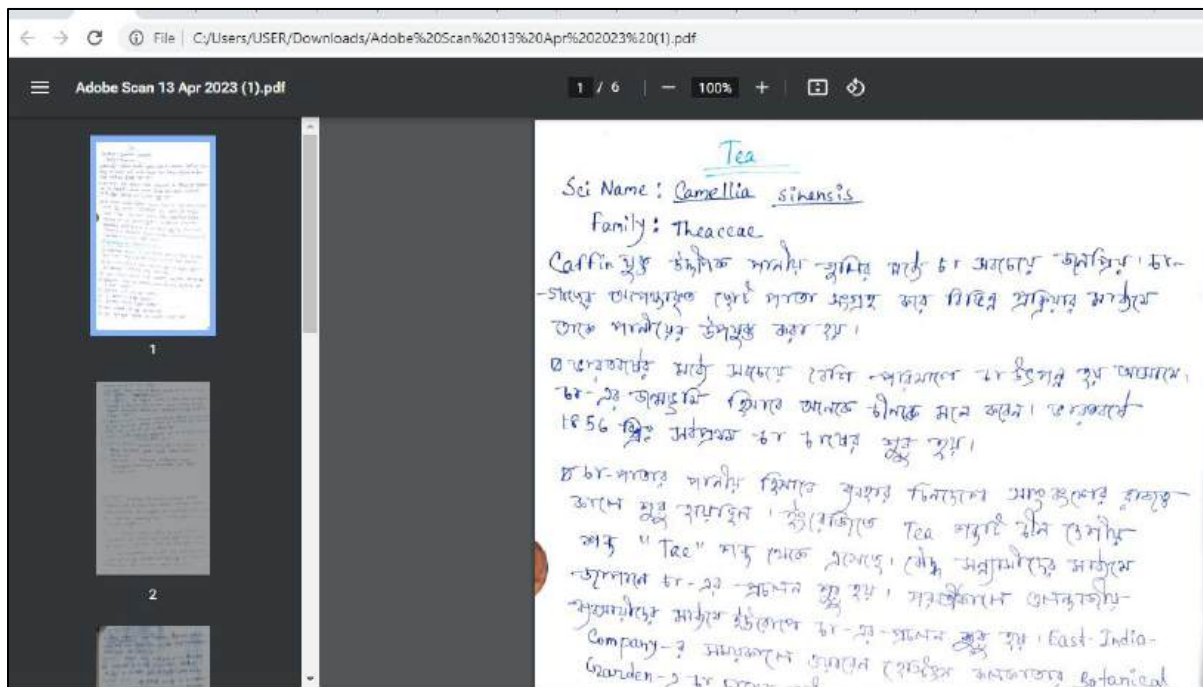
Rehana Hameedulhaque
Chartered Accountant

Word Doc. Sample

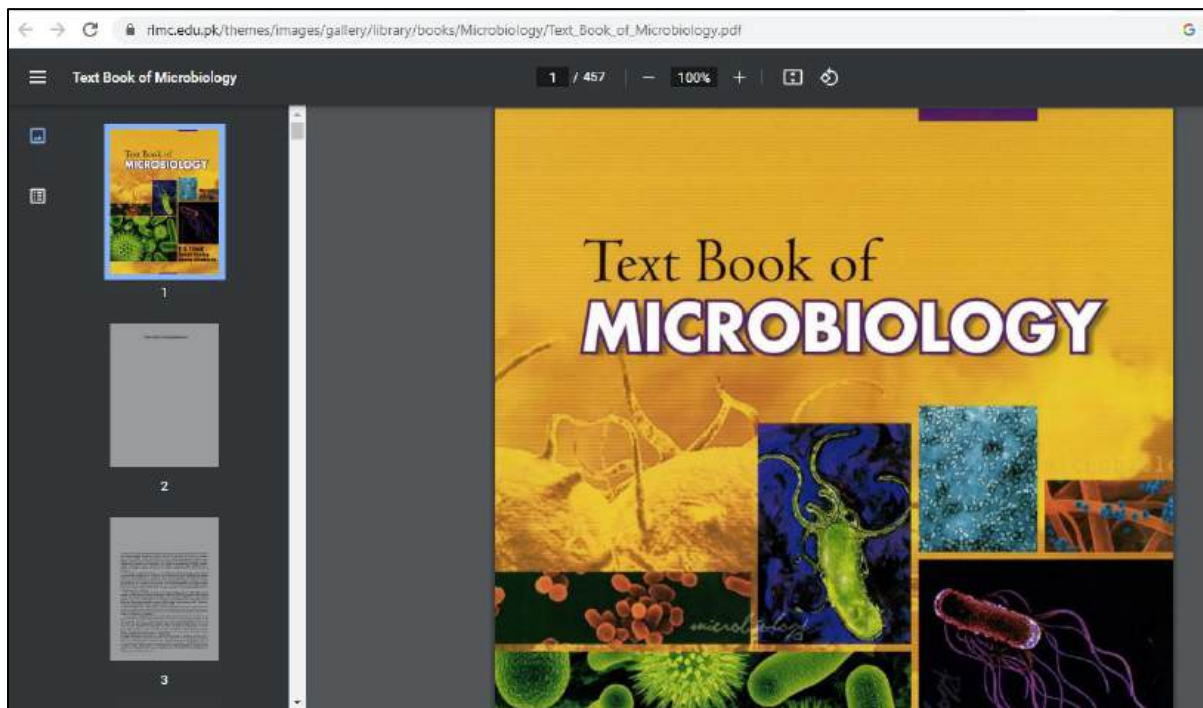


14.08.2024
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Principal, Narayana Group
Channarayana, Mysore

Scanned hand written notes (Sample in Bengali for general stream)



E-Book PDF (Sample)

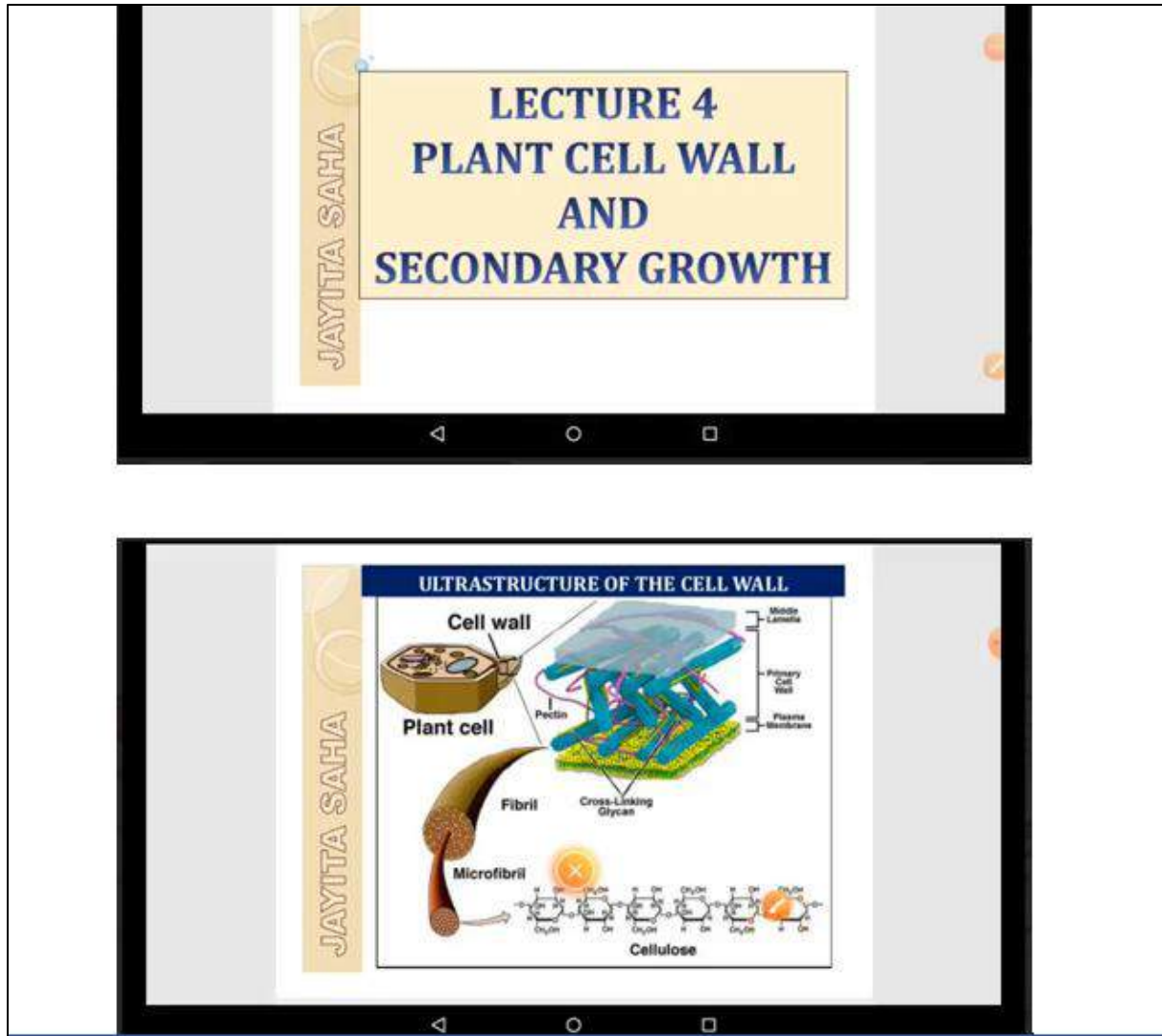


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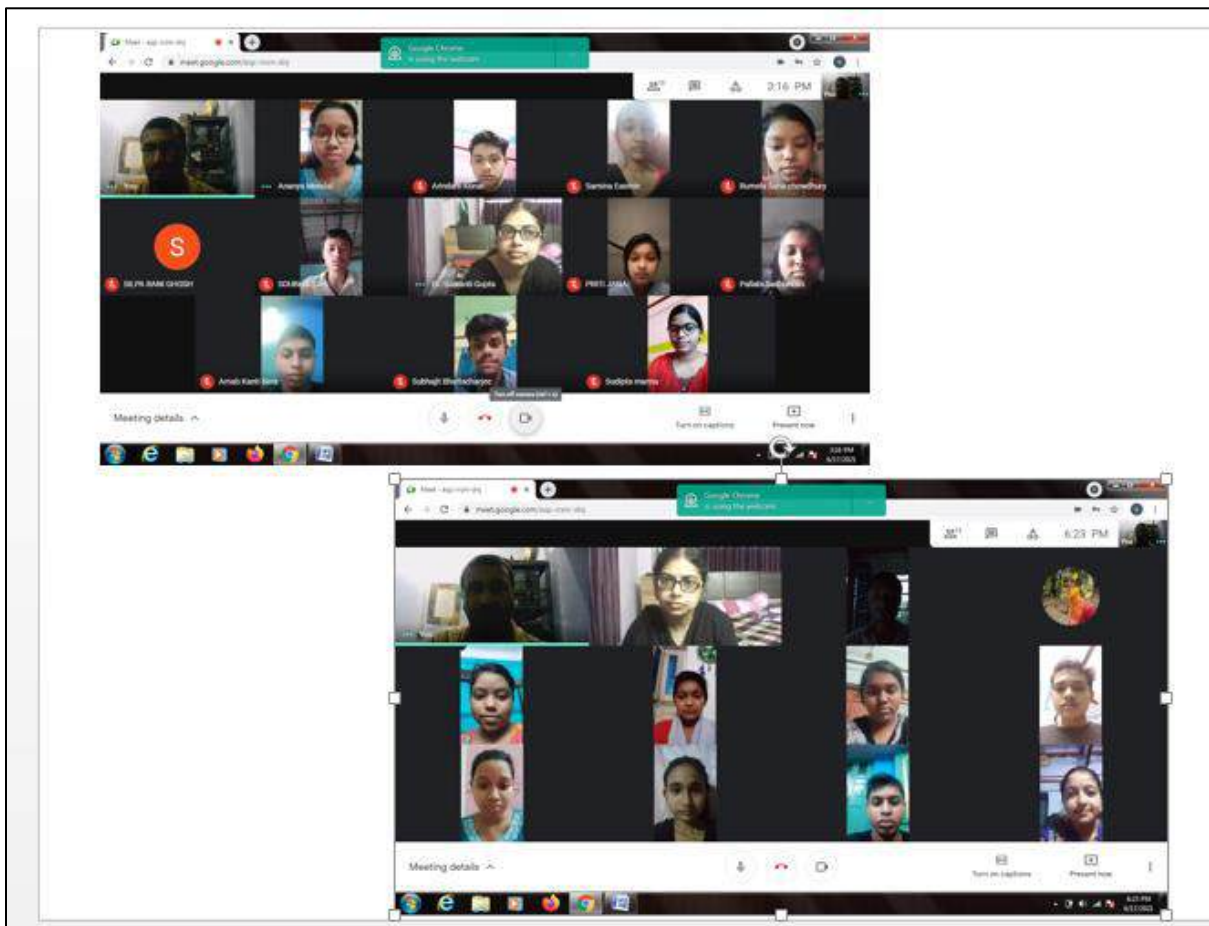
Principal
Rabindranath Mahavidyalaya
Chumprabang, Thongy.

Screenshot of Video (Sample)



14.08.2024
ATTESTED
Principal
Pabna Mahavidyalaya
Chandradhara, Tirohly

Google Meet



14.08.2024
ATTESTED
Principal
Rajivendra Kumar
Chandrasekhar, Tirupur

Department of Chemistry (2021-2022)

Snapshots/ screenshots of E-resources and techniques used

Screenshot (sample) of Resources

PPT Sample

Mechanisms of enzyme inhibition

By
Dr. Sucheta Joy
The Department of Chemistry
Rabindra Mahavidyalaya
Champadanga, Hooghly

1

- Competitive inhibition: the inhibitor (I) binds only to the active site.
 $E + I \rightleftharpoons EI$
- Non-competitive inhibition: binds to a site away from the active site. It can take place on E and ES.
 $E + I \rightleftharpoons EI$
 $ES + I \rightleftharpoons ESI$
- Uncompetitive inhibition: binds to a site of the enzyme that is removed from the active site, but only if the substrate is already present.
 $ES + I \rightleftharpoons ESI$
- The efficiency of the inhibitor (as well as the type of inhibition) can be determined with controlled experiments.

2

3

23.7 Kinetics of photochemical reactions

- Primary photochemical process:** products are formed directly from the excited state of a reactant.
- Secondary photochemical process:** intermediates are formed directly from the excited state of a reactant.
- Photophysical processes compete with the formation of photochemical products via deactivating the excited state.

4

- Times scales of photophysical processes**
Within $10^{-15} \sim 10^{-14}$ s for electronic transitions induced by radiation and thus the upper limit for the rate constant of a first order photochemical reaction is about 10^{15} s^{-1} .
 $10^{-12} \sim 10^{-11}$ s for fluorescence
 $10^{-12} \sim 10^{-11}$ s for intersystem crossing (ISC)
 $10^{-9} \sim 10^{-8}$ s for phosphorescence (large organic molecules)
- A slowly decaying excited species can undergo a very large number of collisions with other reactants before deactivation.
- The interplay between reaction rates and excited state lifetimes** is a very important factor in the determination of the kinetic feasibility of a photochemical process.

5

Autocatalysis

- Autocatalysis: the catalysis of a reaction by its products.
 $A + P \rightarrow 2P$
The rate law is $\frac{d[P]}{dt} = k[A][P]$
- To find the integrated solution for the above differential equation, it is convenient to use the following notations
 $[A] = [A]_0 - x$; $[P] = [P]_0 + x$
- One gets $\frac{d[P]}{dt} = k([A]_0 - x)([P]_0 + x)$
- Integrating the above ODE by using the following relation
gives $\frac{1}{[A]_0 - [P]_0} \ln \frac{[A]_0 + [P]}{[A]([P]_0 + [A])} = k t$
or rearrange into $\frac{1}{[A]_0 - [P]_0} \ln \frac{[A]_0 + [P]}{[A]([P]_0 + [A])} = k t$

14.08.2024
ATTESTED
Principal
Rabindra Mahavidyalaya
Champadanga, Hooghly

POWERPOINT PRESENTATION ON Mrs. Subhra Dholey SACT Department of Chemistry WACKER PROCESS

1

CONTENT

01. Introduction
02. Wacker process
03. Condition and mechanism
04. Product analysis and application
05. Advantages and disadvantages
06. Conclusion

2

CONDITIONS

- Temperature and Pressure**
Reaction is carried through the reactor along with catalyst at 105-110°C and 100-1000 kPa.
- Concentration**
Reaction is carried out in several stages to maintain and control the concentration of catalyst and ligand compounds in catalyst.
- Solvents**
Reaction of ethylene is carried out in aqueous solution. Adding PdCl₂ in the catalyst solution produces a white solid and is added to the catalyst solution. It is added to the catalyst solution after the reaction.

6

STEPS OF THE WACKER PROCESS

- 1) Oxidation of ethylene
- 2) Ligand exchange
- 3) Hydroxy-palladation
- 4) β -hydrogen elimination
- 5) Hydrogen insertion
- 6) Catalyst repressing & product formation

7

ADVANTAGES

- 1. Simplicity: The implementation of the Wacker process is relatively simple and easy to implement.
- 2. Catalyst efficiency: The Wacker process is highly efficient, with a catalyst efficiency of up to 100%.
- 3. High yield: The Wacker process is highly efficient, with a yield of up to 100%.
- 4. Low cost: The Wacker process is highly efficient, with a low cost of operation.

DISADVANTAGES

- 1. Complexity: The implementation of the Wacker process is relatively complex and difficult to implement.
- 2. High cost: The Wacker process is highly efficient, with a high cost of operation.
- 3. Low yield: The Wacker process is highly efficient, with a low yield of up to 100%.
- 4. High waste: The Wacker process is highly efficient, with a high waste of up to 100%.

MECHANISM

POWERPOINT PRESENTATION ON Hydroformylation Reaction Dr. Rabiul Alam Assistant Professor Department of Chemistry

1

CONTENT

- 1. Introduction
- 2. Mechanism
- 3. Applications
- 4. Advantages and Disadvantages
- 5. Conclusion

2

INTRODUCTION

Hydroformylation is a chemical reaction in which an alkene reacts with carbon monoxide and hydrogen to form an aldehyde. This reaction is widely used in the chemical industry for the production of aldehydes, which are then used in the synthesis of various chemicals and polymers.

3

Steps of the hydroformylation

1. Coordination of alkene to the catalyst
2. Insertion of alkene into the metal-hydrogen bond
3. Migratory insertion of CO
4. Hydrogenation of the resulting intermediate

5

MECHANISM

6

USES

- 1. Production of aldehydes for the synthesis of various chemicals and polymers.
- 2. Production of aldehydes for the synthesis of various chemicals and polymers.

7

Conclusion

Hydroformylation is a chemical reaction in which an alkene reacts with carbon monoxide and hydrogen to form an aldehyde. This reaction is widely used in the chemical industry for the production of aldehydes, which are then used in the synthesis of various chemicals and polymers.


PRODUCTS OF HYDROFORMYLATION REACTION

MECHANISM OF HYDROFORMYLATION REACTION

14.08.2024
ATTESTED
Rabiul Alam
Assistant Professor
Department of Chemistry, Tatyasaheb Kore

Chemistry Hons. SEM- VI
PHYSICAL CHEMISTRY
Molecular Spectroscopy
By
Dr. Debasmita Sarder
Assistant Professor
Department of Chemistry
Rabindra Mahavidyalaya
Chunaganj, Hooghly

Course Code: CC-14
Course Title: Physical Chemistry-IV (Thao)
Molecular Spectroscopy



Topics

- 1. Introduction to X-ray spectroscopy: radiation with electrons and various types of X-ray tubes. Capabilities & applications
- 2. Absorption spectroscopy: Selection rules, intensities of spectral lines, determination of band lengths of elements and trace elements. Methods, sample preparation
- 3. Fluorescent spectroscopy: Basic principles of radiation; excitation of their various, analysis of elements, molecular substances, pharmaceuticals. Linear calibration, detection limits, background, interferences, quenching, scattering, loss bands, analysis of elements for polyatomic molecules, mode of detection
- 4. Raman spectroscopy: Qualitative treatment of Raman effect (Effect of incident photon, vibrational frequency, Stokes and Anti Stokes lines) their intensity distribution, use of medical analysis
- 5. Nuclear Magnetic Resonance (NMR) spectroscopy: Principles of NMR spectroscopy, terms precession, chemical shift, spin-spin coupling, different nuclei, spin-spin coupling and high-resolution NMR

The diagram illustrates the propagation of electromagnetic radiation. A wave is shown moving to the right, indicated by an arrow labeled "Direction". The wave consists of a blue sinusoidal curve representing the "Magnetic Field" and a green sinusoidal curve representing the "Electric Field". The two fields are perpendicular to each other. A "Source" is indicated at the beginning of the wave. A legend on the right shows a blue arrow for "Magnetic Field" and a red arrow for "Electric Field", both pointing in the "Direction" of propagation.

Regions of Spectrum
THE ELECTROMAGNETIC SPECTRUM

Radio waves Microwaves Infrared radiation Visible light Ultraviolet X-rays Gamma rays

What are the typical electromagnetic radiation?

Regions of Electromagnetic Radiation

Radio	TV	FM	AM	Visible	UV	X-ray	Gamma
Wavelength	1000 m	100 m	10 m	1000 nm	100 nm	10 nm	10 pm
Frequency	100 MHz	100 MHz	1000 kHz	100 THz	100 PHz	100 EHz	100 ZHz

Visible spectrum colors: Red, Orange, Yellow, Green, Blue, Violet.

Gamma ray: γ

Diagram labels: Radio, TV, FM, AM, Visible, UV, X-ray, Gamma.

PRINCIPLE

- When macroeconomic indicators are included upon a sample then this signal resonates with the sample in some fashion. It can be understood as captured in some manner. It is the meaning of the indicator that allows which gives information about molecular structure.
- Resonance is based on measuring. This sample is identified with a selected source, typically a laser. Most of the indicator is captured within the selected source.
- A small portion is taken out of the sample and is used to create a small portion of the sample (such as a laser beam). This laser beam is what we use to identify the sample. It is the meaning of the indicator that allows which gives information about molecular structure.

Anharmonicity

$$E_v = \omega_e(v + 1/2) - \omega_e x_e(v + 1/2)^2 + G(v + 1/2)^3 + \dots$$

Selection Rules:
 $\Delta v = \pm 1, \pm 2, \dots$
 $\Delta v = \pm 2, \pm 3, \dots$ are called overtones.

Dyphonics are often weak because anharmonicity of low v is small.

Figure 19-10. Potential energy diagram for an anharmonic oscillator. The potential energy curve is shown as a function of internuclear distance. The vibrational energy levels are indicated by horizontal lines. The ground state is labeled $v=0$. The first overtone transition is labeled $v=0$ to $v=2$. The first overtone transition is labeled $v=0$ to $v=2$. The first overtone transition is labeled $v=0$ to $v=2$.

POWER POINT PRESENTATION

"WILKINSON'S CATALYST"

Mr. Tanmoy Pandit

SACIT
Department of Chemistry
Rutisha Mahavidyalaya
Churupsinga, Hooghly

1.1. THE SCIENCE OF CHEMISTRY

- INTRODUCTION
- ORGANIC CHEMISTRY: CARBON
- INORGANIC AND PHYSICAL
- BIOCHEMISTRY
- TRANSCHEMISTRY
- UNIT
- APPLIED CHEMISTRY
- ADVANCED CHEMISTRY
- ORGANOMETALLIC
- ORGANOPHOSPHORUS

ON MODERN SCIENCE

- A catalyst is a substance that affects the rate of reaction but remains from the process unchanged.
- A catalyst usually changes a reaction via its intermediate level instead path (lower energy) for the reaction.
- Catalysts work both reversible and irreversible.
- Change only the rate of reaction, it does not affect thermodynamic.

Wilkinson's catalyst is named after the English chemist and Nobel laureate Sir Geoffrey Wilkinson.

Wilkinson's catalyst is a homogeneous catalyst consisting of a rhodium complex with a triphosphine ligand and a hydride ligand.

where $\text{P}(\text{C}_6\text{H}_5)_3$ is triphenylphosphine.

It is extremely useful as a catalyst in the hydrogenation of alkenes.

The structure is given below:

At ambient temperatures, this complex is a pink solid in a neutral, square planar geometry.

Wilkinson's Catalyst

[illegible][illegible]

12

Reaction scheme showing the synthesis of 1,2,3,4-tetrahydro-6-methyl-5H-benzothiazine-5-thione (12) from 2-mercapto-1-methyl-4-methyl-5H-benzothiazine-5-thione (11) and 2-mercapto-1-methyl-4-methyl-5H-benzothiazine-5-thione (10). The reaction is catalyzed by CuI in DMF at 100°C for 24 hours.

APPLICATIONS

- **hydrophobicity** (power of accumulation) hydrophobic: "the chemical nature of substances, which impedes hydrophobic molecules from the solution of polarized solvent like water"
 - in the cellular "lipid bilayer" going to affect the different parts can be critical
- it is very important in the cellular hydrophobic "process"
- in the cellular hydrophobicity of "alliance" (the chemical nature of substances, which impedes hydrophobic molecules from the solution of polarized solvent like water)



Advantages	Disadvantages
<ul style="list-style-type: none"> Relatively high sensitivity Relatively low reagent consumption Can use specific, selective, and directed reagents for analysis Low reagent cost Generally, low reagent addition to single analysis 	<ul style="list-style-type: none"> Complex reagent preparation Complex reagent transportation and storage Complex reagent storage Complex reagent disposal Reagent waste High reagent cost High reagent consumption Low selectivity Low sensitivity Low reagent stability

14.08.2024

ATTESTED
Principal
Rajniya Narayanaswamy
Channarayana, Hoysaly

PDF Sample/ Books

TERPENOIDS


* Terpenes \Rightarrow Terpenes are naturally occurring hydrocarbons having a structural relationship with isoprene, i.e. 2-methyl-1,3-butadiene ($\text{H}_2\text{C}=\text{C}(\text{CH}_3)-\text{CH}=\text{CH}_2$)

3420
115-1
Iso prene, itself, doesn't occur in nature.

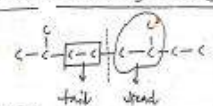
* Terpenoids \Rightarrow Terpenoids include hydrocarbons as well as their oxygenated derivatives, such as alcohols, aldehydes, ketones etc.

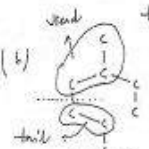
Most of the terpenoids are unsaturated aliphatic or cyclic compounds.

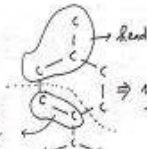
* Iso-prene rule \Rightarrow Thermal decomposition of almost all terpenoids produces iso-prene as one of the products and the skeleton structure of all naturally occurring terpenoids can be constructed by joining iso-prene units as required following head to tail connection.

$\text{H}_2\text{C}=\text{C}(\text{CH}_3)-\text{CH}=\text{CH}_2$ / structural skeleton \Rightarrow 

* Head to tail connectivity in the following \Rightarrow

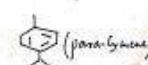
(a) 

(b) 

(c) 

(b) (cyclic structure) / (acyclic mono-terpene)

(c) (para-cymene structure)

 (para-cymene)

4.08.2024

ATTESTED


Principal
Pabna Mahavidyalaya
Chattogram, Pabna

Biocatalysis for Green Chemistry and Chemical Process Development

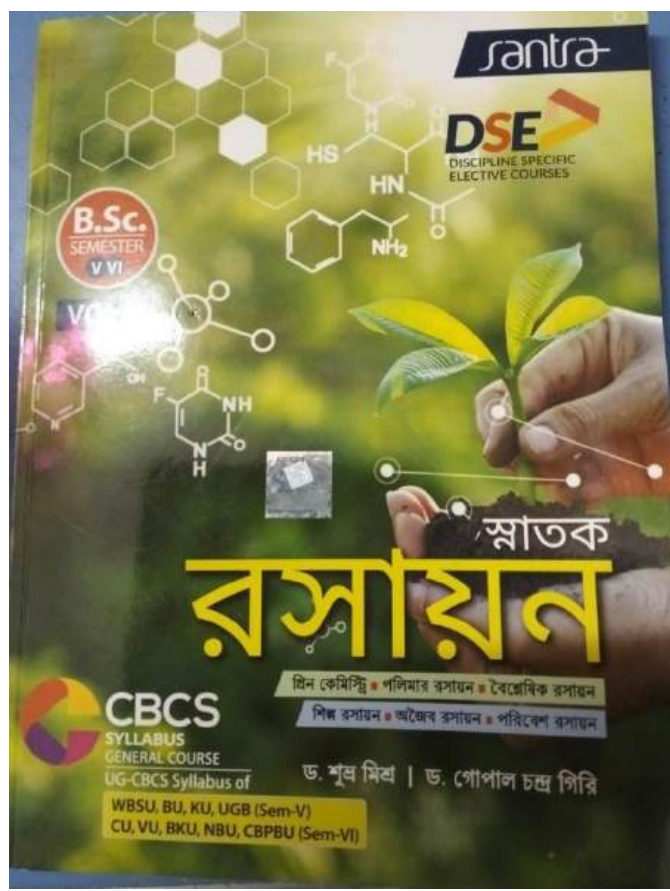


Edited by
Junhua (Alex) Tao
Romas Kazlauskas

 WILEY

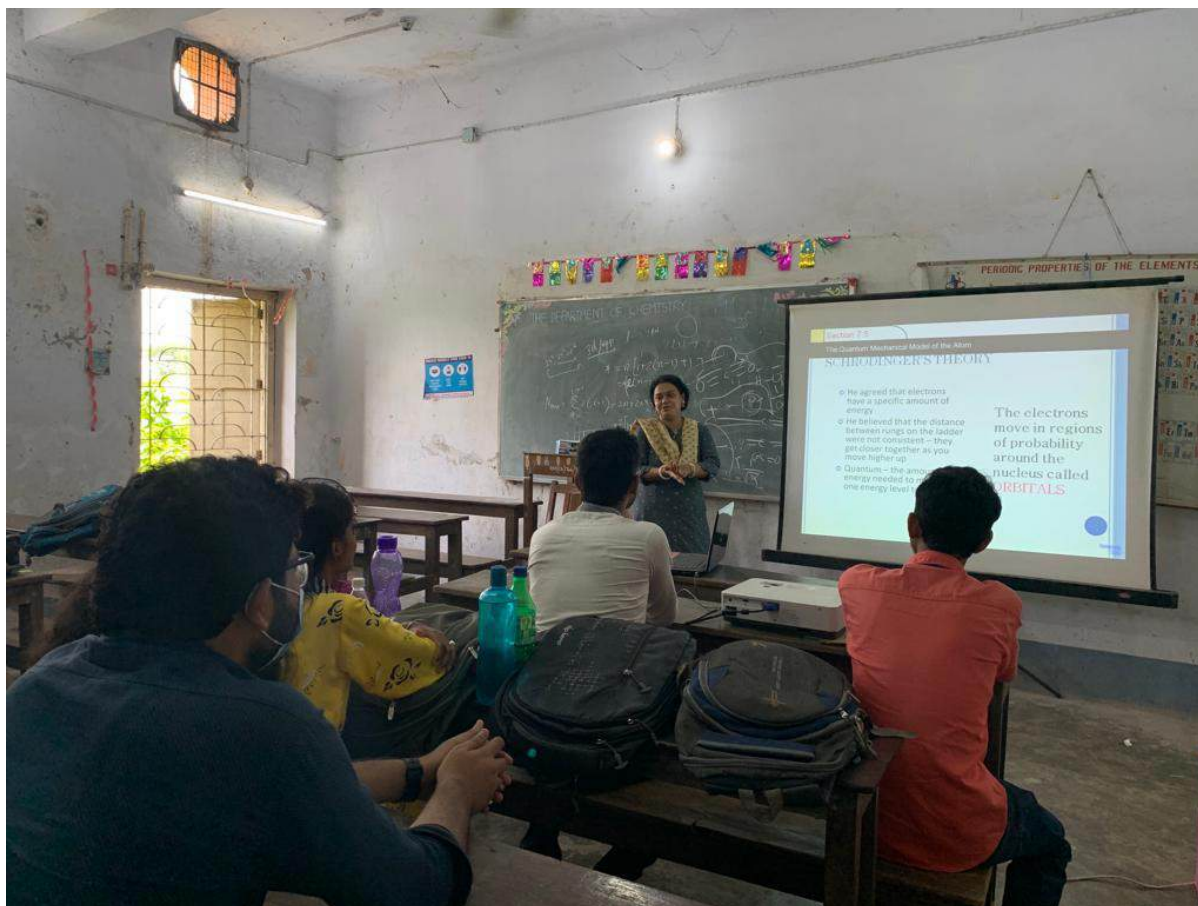

14.08.2024
ATTESTED
Principal
Referee Methodology
Chempeddy, Tirogally

Bengali book for general students



14.08.2024
ATTESTED
Principal
Rabindra Mahavidyalaya
Chandrapur, Tinsukia

ICT Classes



Class taken by Dr. Sucheta Joy, Assistant Professor, department of Chemistry

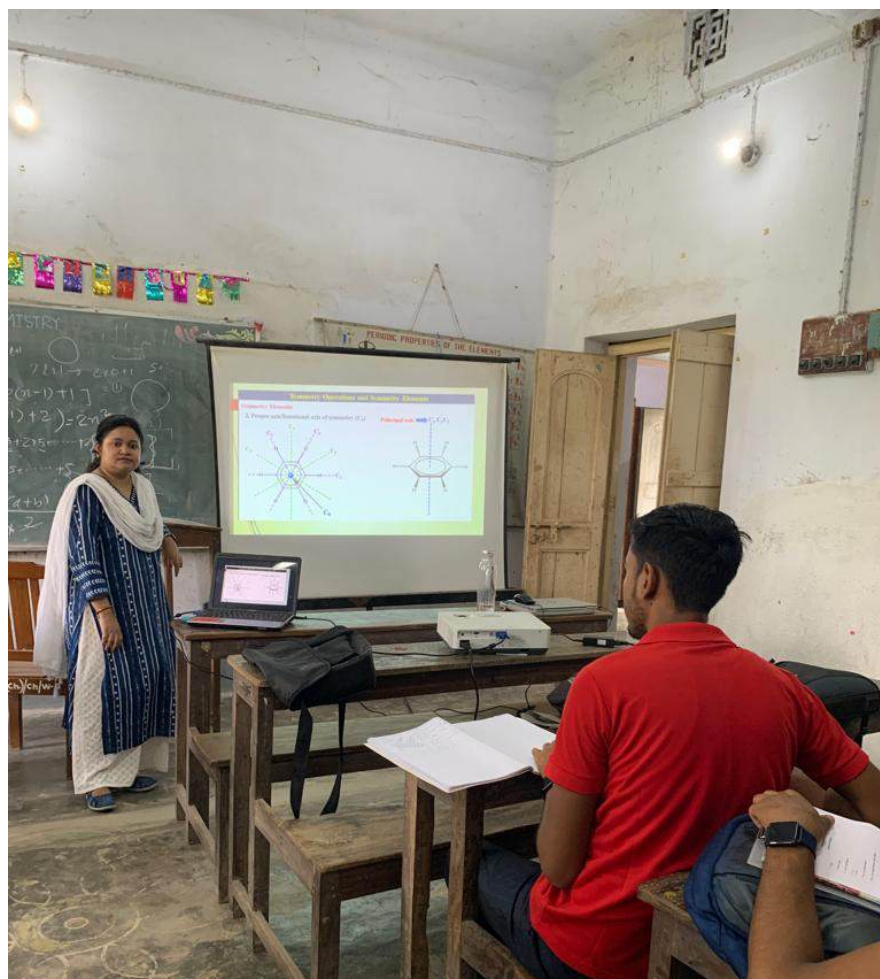
14.08.2024
ATTESTED
Principal
Rachina Maheshwari
Chandrasekhar, Tirunelveli




14.08.2024
ATTESTED
Principal
Rajendra Maheshwari
Chandrapur, Nagpur



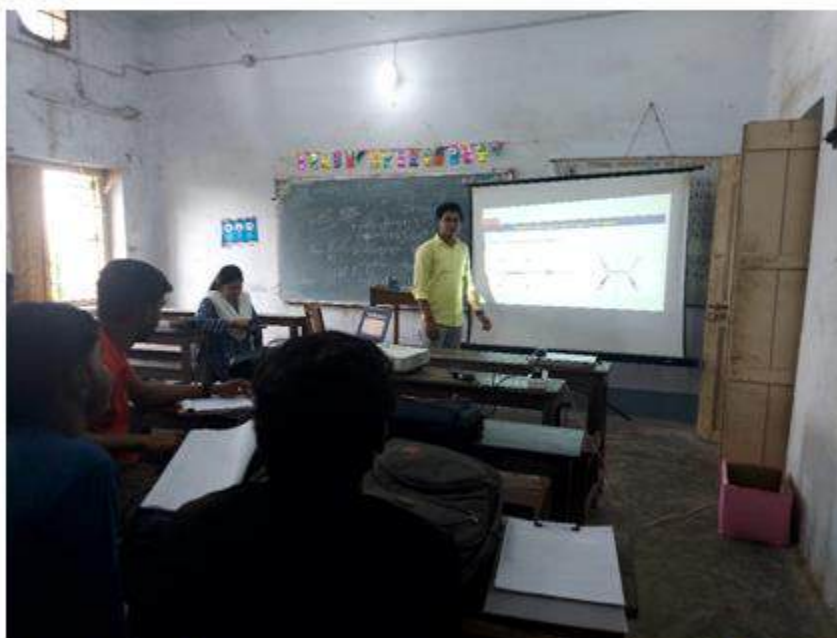
Class taken by Dr. Rabiul Alam, Assistant Professor, department of Chemistry





 14.08.2024
 ATTESTED
 Principal
 Pabna Mahavidyalaya
 Chaudhury, Pabna



Class taken by Dr. Debasmita Sardar, Assistant Professor, department of Chemistry



Class taken by Mr. Tanmoy Pandit, SACT, department of Chemistry

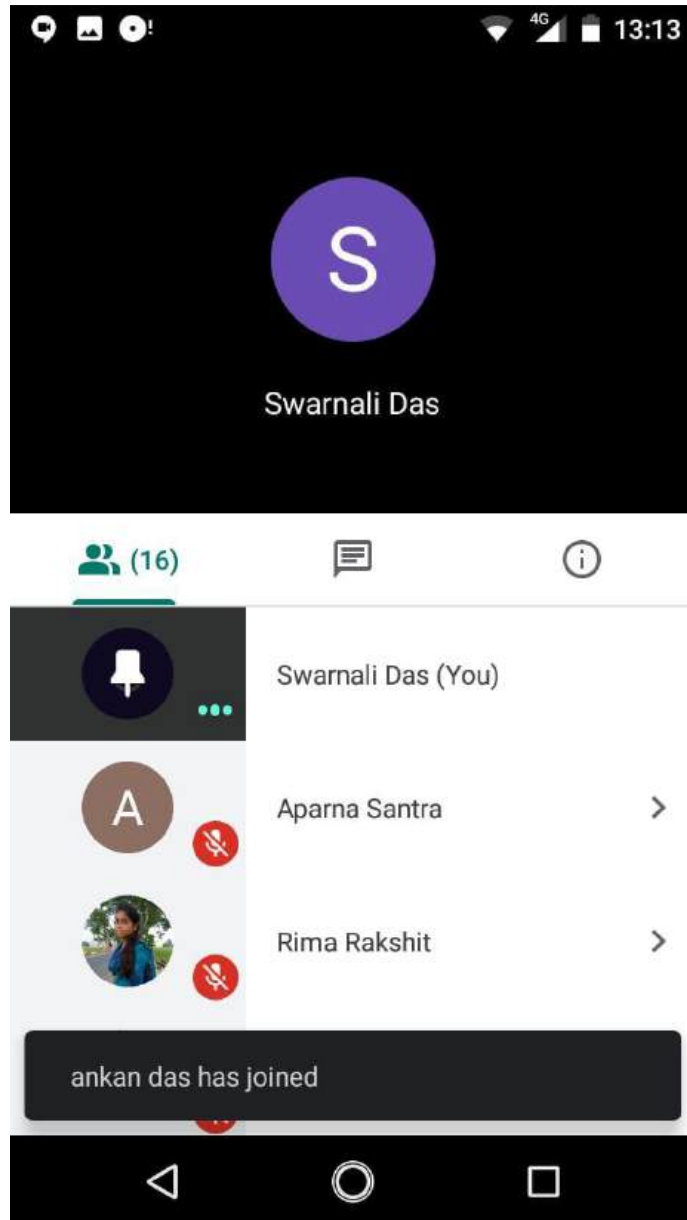

 14.08.2024
 ATTESTED
 Principal
 Pabarna Mahavidyalaya
 Champadanga, Tangaila



Class taken by Mrs. Subhra Dholey, SACT, department of Chemistry

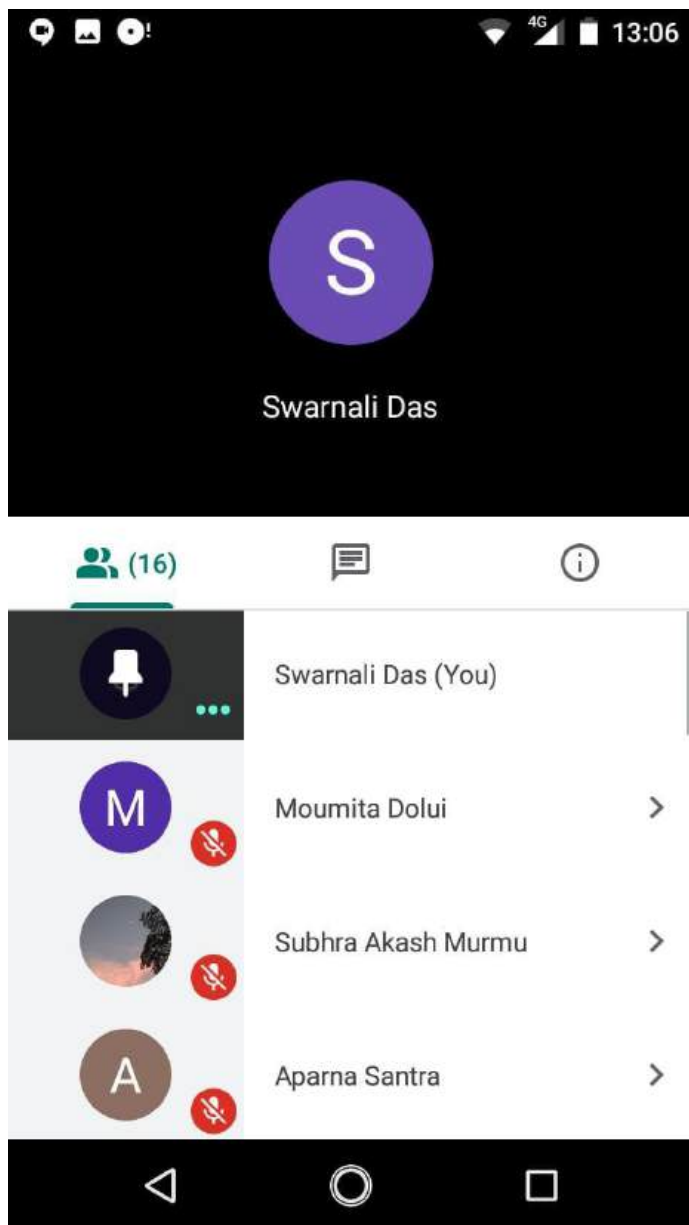
14.08.2024
ATTESTED
Principal/
Rishika Maheshwari
Chemistry, Tatyga

Department of Political Science



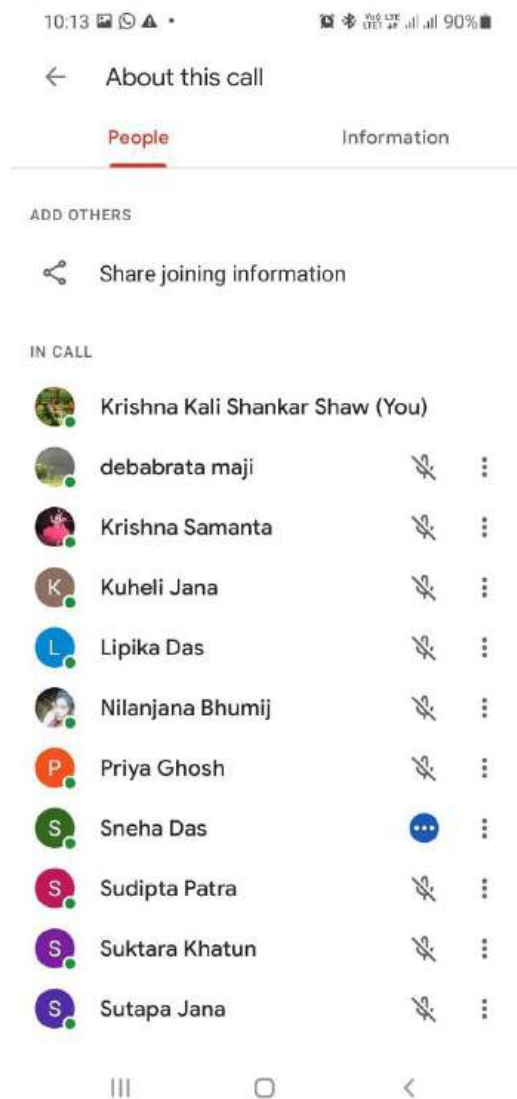
Online class taken by Swarnali Das (SACT) for 4 th sem generic

14.08.2024
ATTESTED
Principal
Rabindra Mahavidyalaya
Chandannagar, Hooghly



Online class taken by Swarnali Das SACT for Sem-6 Generic

14.08.2024
ATTESTED
Principal
Rajendra Kumar Choudhary
Chandigarh, Haryana



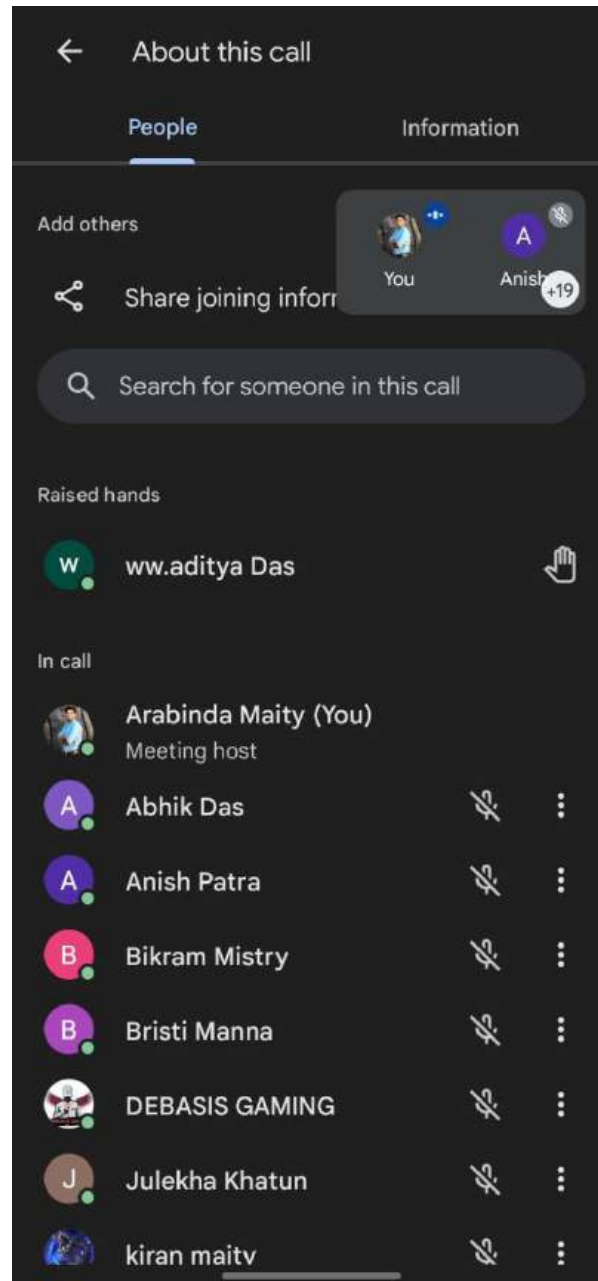
Online Class taken by Mr. Krishna Kali Mondal Asst. Prof in Pol. Sc for 2 nd Sem Honours

14.08.2024
 ATTESTED
 Principal
 Pabna Mahavidyalaya
 Chaudhary, Firozpur

DEPARTMENT OF PHYSICAL EDUCATION (SESSION 2021-2022)

ONLINE CLASS - SEMESTER-1

CLASS TAKEN BY MR. ARABINDA MAITY , SACT, PHYSICAL EDUCATION



Rests of the departments also have followed the similar techniques to conduct classes and evaluate student performance. (Screenshots are not attached.)

14.08.2024
ATTESTED
Principal
Chandrapada, Hooghly

Session 2020-21

Teachers use ICT enabled tools for effective teaching-learning process.


14.08.2024
ATTESTED
Principal
Rajwade Mahavidyalaya
Chandrapur, Hoshiy.



RABINDRA MAHAVIDYALAYA

Affiliated to the University of Burdwan
Champadanga, Hooghly, West Bengal, Pin.-712401


Estd.- 1971




Ref. No.....

Date.....

This is to certify that the following ICT enabled tools for effective teaching-learning process was used by the various Departments in the session 2020-2021 at Rabindra Mahavidyalaya, Champadanga, Hooghly, West Bengal


Principal
Rabindra Mahavidyalaya
Champadanga - Hooghly

Dr. Prasanta Bhattacharyya
Principal


14.05.2024
ATTESTED
Principal
Rabindra Mahavidyalaya
Champadanga, Hooghly

Session 2020-21

Teachers use ICT enabled tools for effective teaching-learning process.

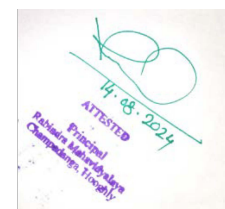
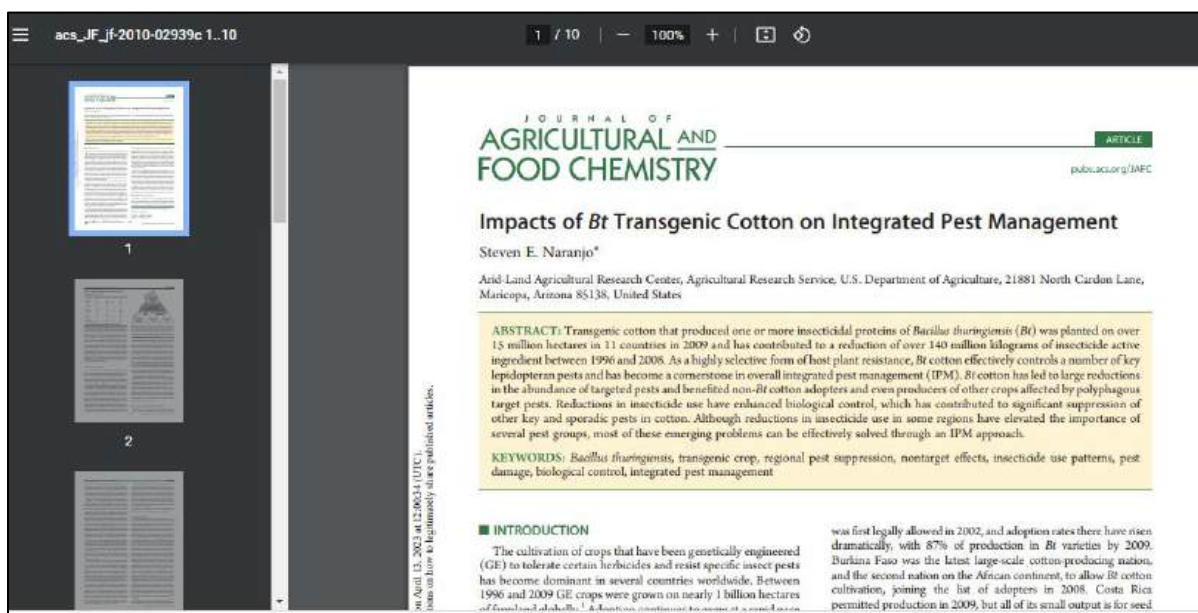
Due to Covid-19 entire teaching learning was switched to online mode. Google Meet, Zoom, etc were used for teaching through. desktop, Laptop. Smart Phone etc. Video, PPT, PDF, excel, etc files were shared for study material. E-Journals and e-books were used for effective teaching and learning.

Some screenshots of the same are attached here.

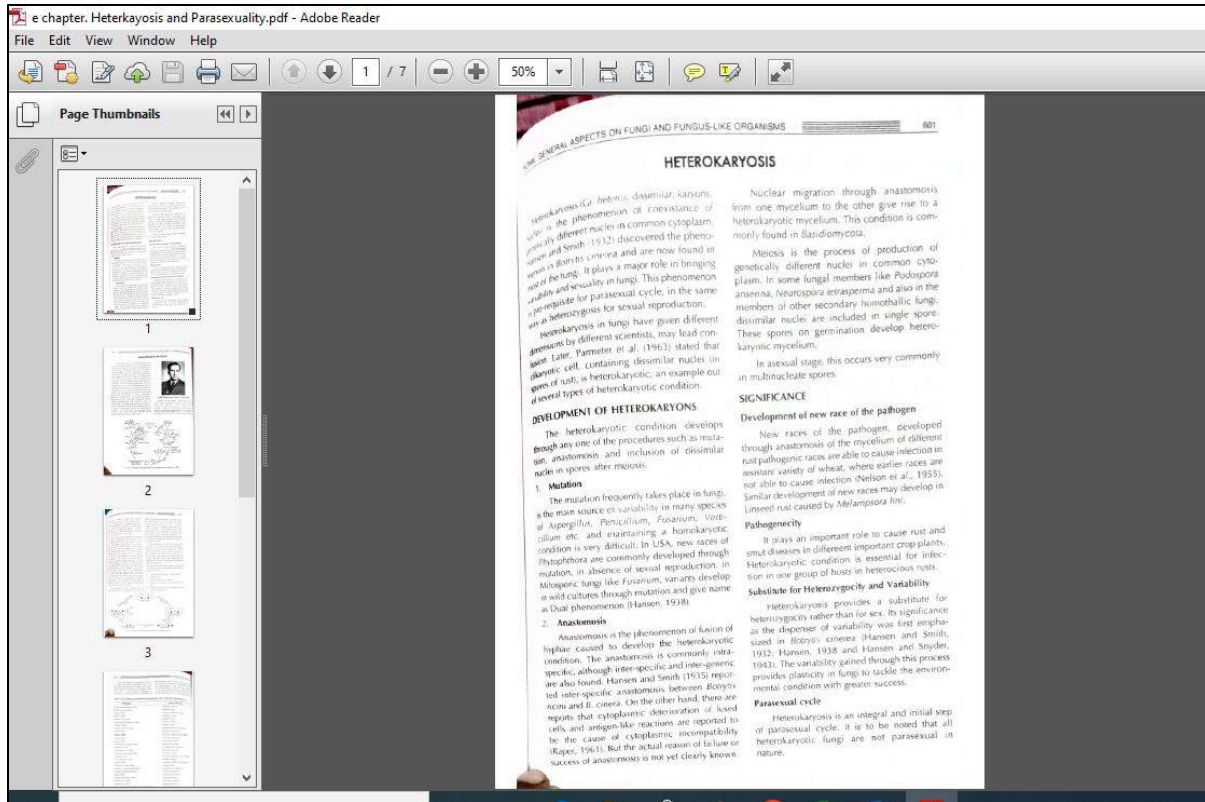
Department of Botany (2020-2021)

Snapshots/ screenshots of E-resources and techniques used

Screenshot (sample) of E- Journal Resource



PDF Sample



14.08.2024
ATTESTED
Principal
Fakir Mohan Mahavidyalaya
Chandannagar, Hooghly

PPT. Sample

INDUCIBLE AND REPRESSIBLE OPERON: POSITIVE AND NEGATIVE CONTROL

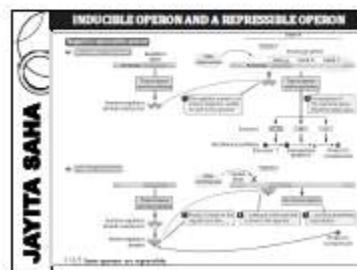
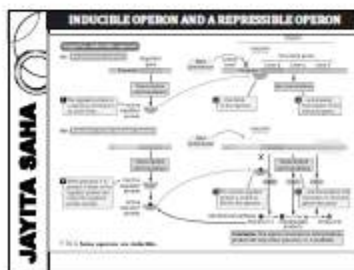
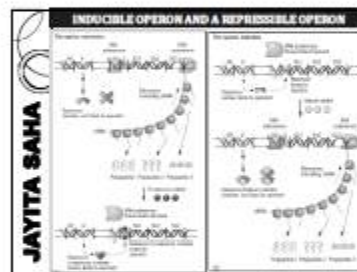
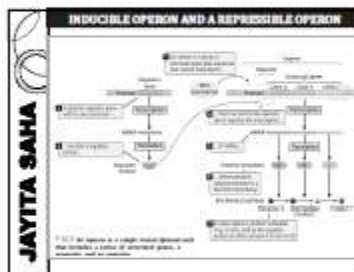
INDUCIBLE OPERON AND A REPRESSIBLE OPERON

INDUCIBLE OPERON:

- The repressor binds to the operator(s) of an inducible operon.
- Inducible operon is turned off in the absence of the effector (inducer) molecule.
- Lac operon.

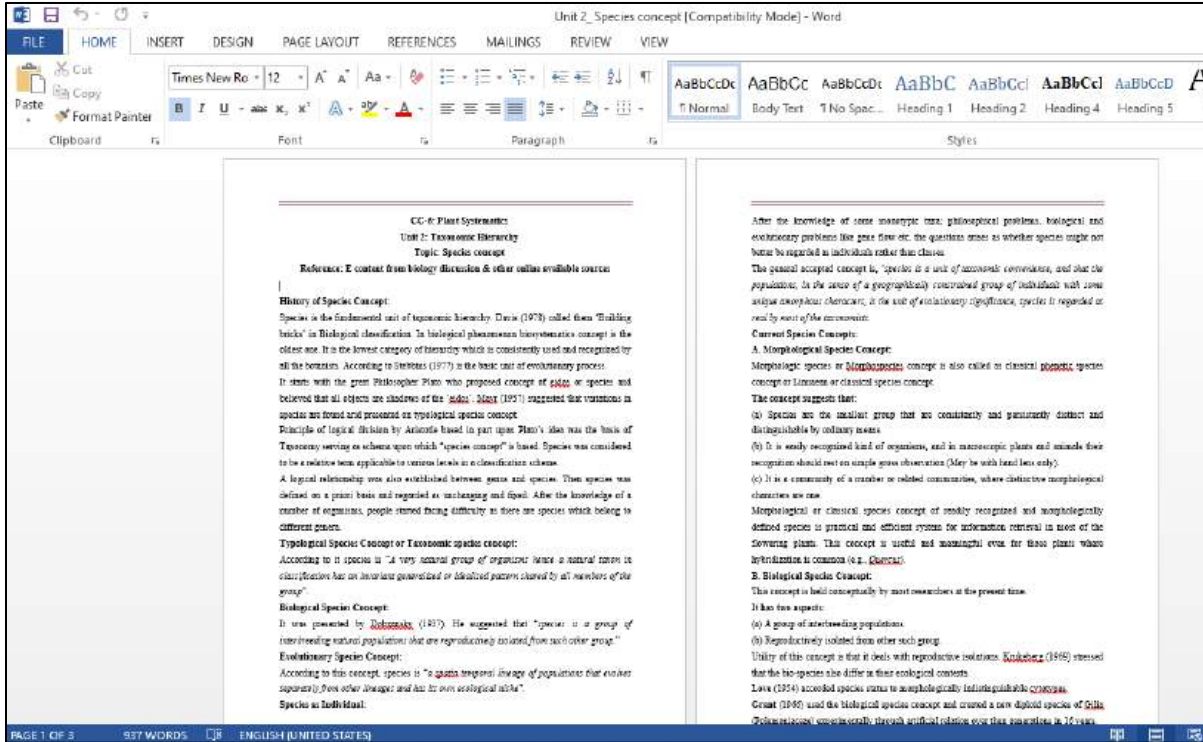
REPRESSIBLE OPERON:

- The repressor/effector molecule complex binds to the operator(s) of a repressible operon.
- A repressible operon is turned on in the absence of the effector (corepressor) molecule.
- Trp operon.



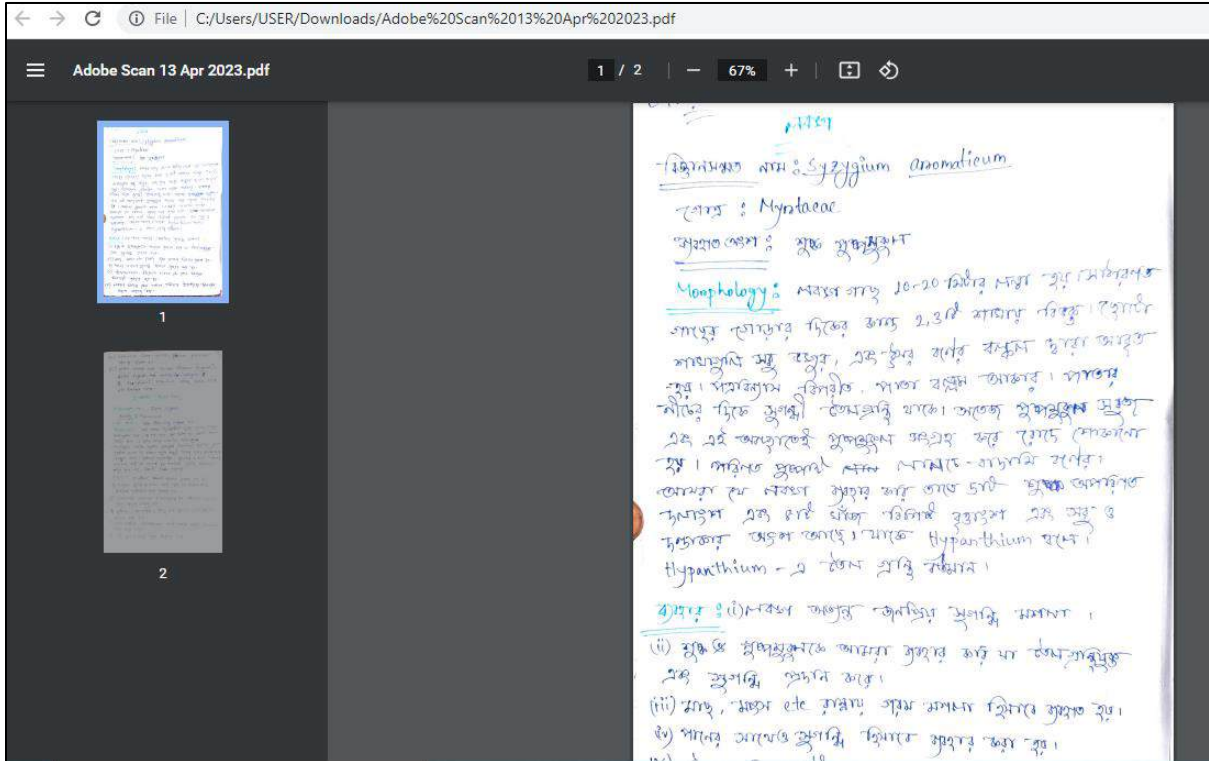
14.08.2024
ATTESTED
Principal
Rabindra Mahapatra
Chandrapur, Bhopal

Word Doc. Sample



14.08.2024
ATTESTED
Principal
Rabindra Kumar
Chatterjee, Trapaty

Scanned hand written notes (Sample in Bengali for general stream)

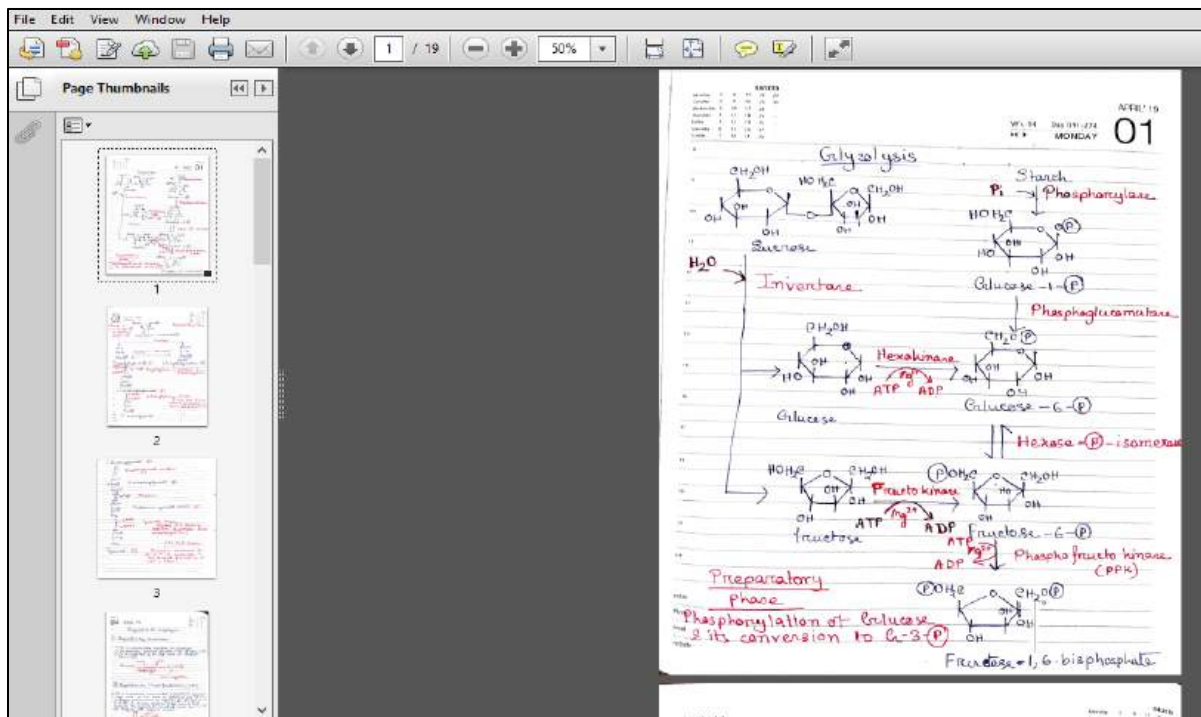


14.08.2024

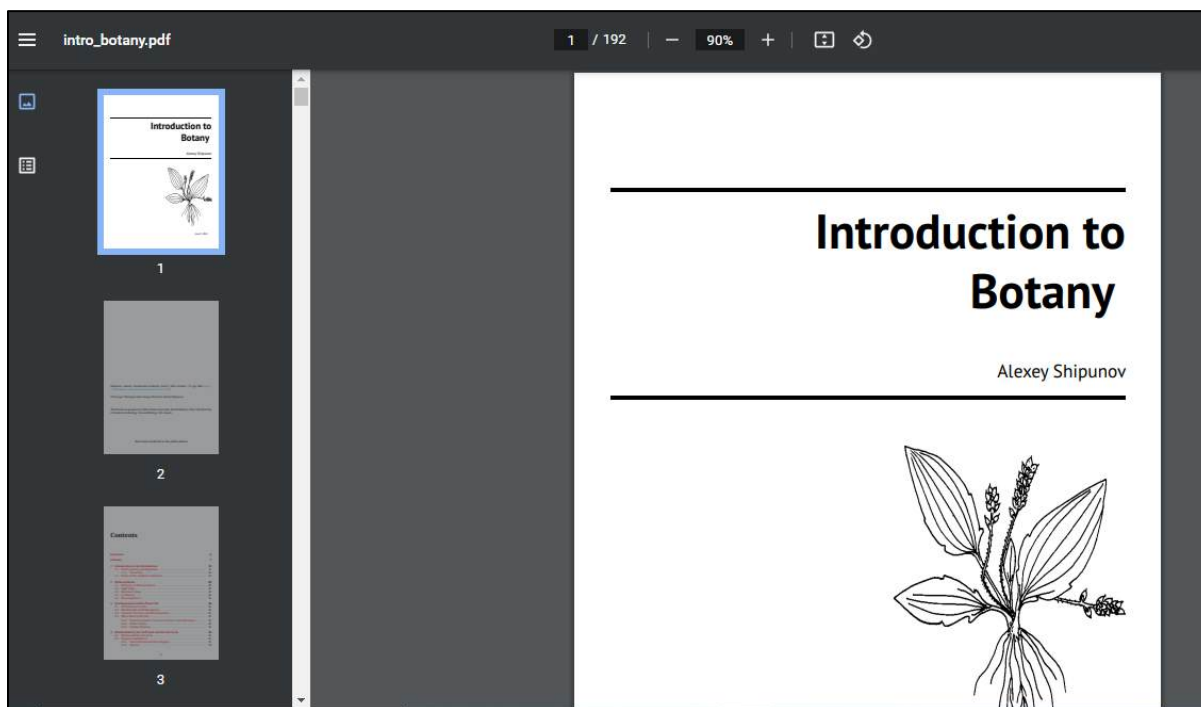
ATTESTED

Prabir Nath
Prabir Nath
Chandragiri, Hooghly

Scanned hand written notes (Sample for Honours stream)



E-Book PDF (Sample)



14-08-2024
ATTESTED
Principal
Rajesh Kumar Singh
Chandigarh University

Google Classroom

The screenshot shows the Google Classroom interface for a class titled "SEMESTER IV AND V HONOURS AND GENERAL". The URL in the address bar is classroom.google.com/u/1/c/MTU3OTgzOTk3MjA4. The top navigation bar includes "Stream", "Classwork", "People", and "Marks". The main content area features a header banner with a green background and a globe, with a "Customise" button. Below the banner, the class code is "sjlynd5". There is an "Upcoming" section indicating "No work due in soon" with a "View all" link. The "Stream" section shows a post from "RABINDRA MAHAVIDYALAYA" dated 25 Nov 2020, titled "posted a new assignment: Only for General Students". Below this, there are two PDF assignments titled "Cellular and Molecular Im..." with a "PDF" icon. At the bottom, there is another post from "RABINDRA MAHAVIDYALAYA" dated 25 Nov 2020, titled "posted a new assignment: Only for Honours Students".

The screenshot shows the Google Classroom interface for a class titled "SEMESTER II AND III HONOURS AND GENERAL". The top navigation bar includes "Stream", "Classwork", "People", and "Marks". The main content area features a header banner with a green background and a DNA helix, with a "Customise" button. Below the banner, the class code is "6qgeg4b". There is an "Upcoming" section indicating "No work due in soon" with a "View all" link. The "Stream" section shows a post from "RABINDRA MAHAVIDYALAYA" dated 25 Nov 2020, titled "posted a new assignment: Only for General Students". Below this, there is another post from "RABINDRA MAHAVIDYALAYA" dated 25 Nov 2020, titled "posted a new assignment: Only for Honours Students". At the bottom, there is a third post from "RABINDRA MAHAVIDYALAYA" dated 9 Nov 2020, titled "posted a new assignment: Only for General Students".

14-08-2024
ATTESTED
Principal
Rabindra Mahavidyalaya
Chandigarh, Punjab

Covid -19 Pandemic: Overview

Stream Classwork People Grades

Covid -19 Pandemic: Overview

Class code: **qtqqd45**

Upcoming
No work due soon
[View all](#)

Dr. Jayita Saha posted a new material: Class Material on "COVID-19 PANDEMIC: OVERVIEW"

Posted Aug 4, 2020
You have to read the material before doing the assignment.

Drive file
Unknown File

7 class comments

SEMESTER I and II (2020-21)
HONOURS AND GENERAL

Stream Classwork People Grades

SEMESTER I and II (2020-21)
HONOURS AND GENERAL

Class code: **74ej6dm**

Upcoming
No work due soon
[View all](#)

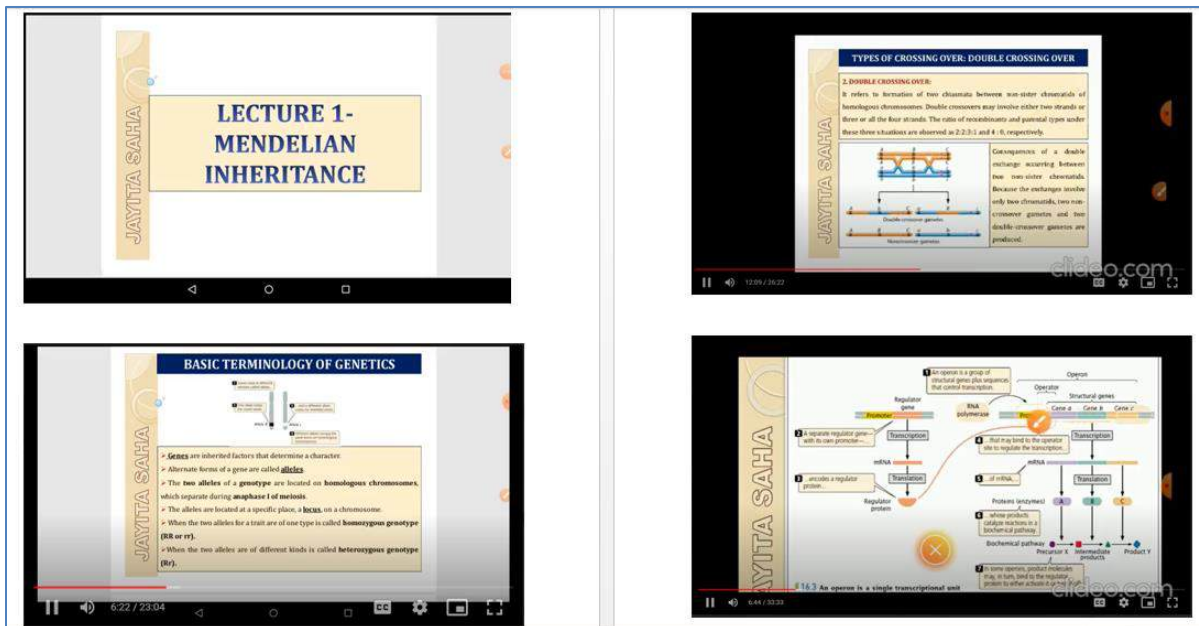
Dr. Jayita Saha posted a new assignment: SEM II [G] : Submission of Assignment
Jul 22, 2021 (Edited Aug 25, 2021)

Dr. Jayita Saha posted a new assignment: SEM II [H] : Submission of Assignment
Jul 22, 2021 (Edited Aug 25, 2021)

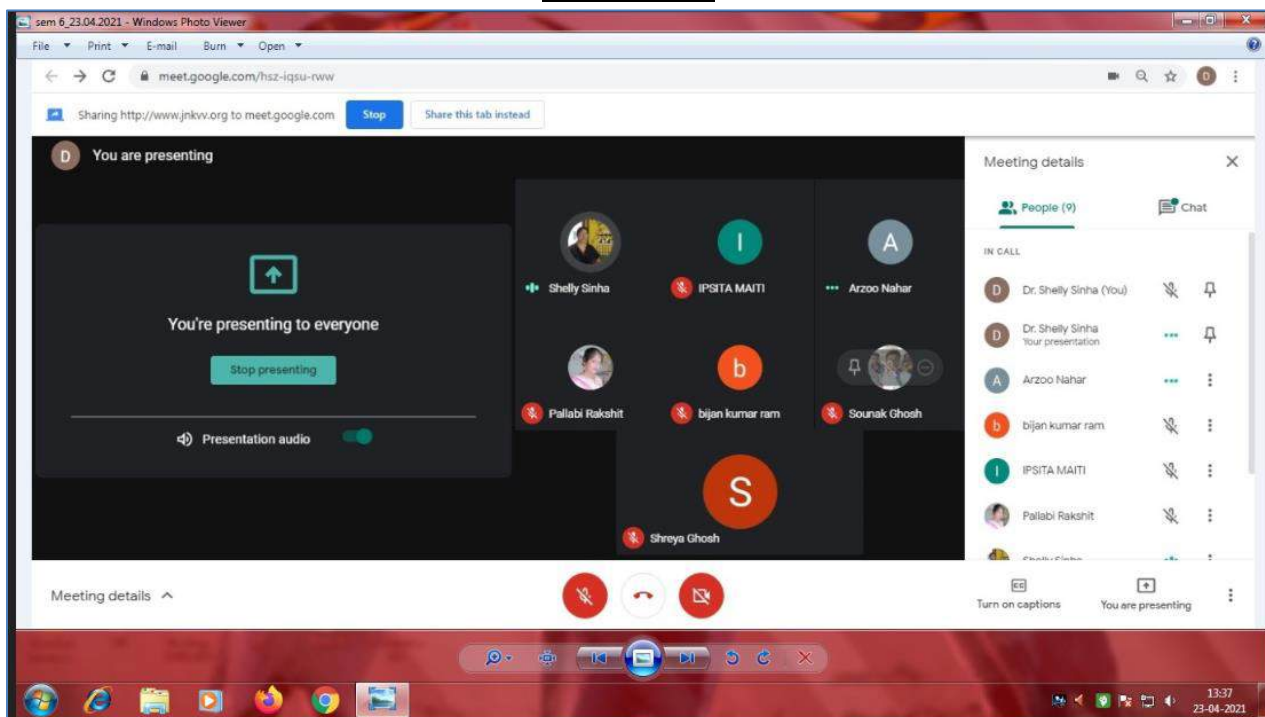
Dr. Jayita Saha posted a new assignment: SEM II [G] : Submission of Assignment
Jun 8, 2021 (Edited Aug 25, 2021)

14.08.2024
ATTESTED
Principal
Rahana Mahabubul Alam
Chandannagar, Hooghly.

Screenshot of Video (Sample)



Google Meet



14.08.2024
 ATTESTED
 Principal
 Rajwade Mahavidyalaya
 Chhapra, Patna

Department of Chemistry (2020-2021)

Snapshots/ screenshots of E-resources and techniques used

Screenshot (sample) of Resources

PPT Sample



14.08.2024
ATTESTED
Principal
Rajni Kant Sharma
Chandigarh, India

Chemistry Hons. SEM-IV
PHYSICAL CHEMISTRY
Applications of Thermodynamics-II
 By
Dr. Debasmita Sardar
 Assistant Professor
 Department of Chemistry
 Rabindra Mahavidyalaya
 Champdanga, Hooghly

19.03.2021

10

Course Code- CC-8
Course Title- Physical Chemistry-III (Theo)
Applications of Thermodynamics-II

11

Topics
Application of Thermodynamics - II

1. Colligative properties: Vapour pressure of solution; Ideal solutions; Ideal dilute solutions and colligative properties; Raoult's law; Thermodynamic derivation using chemical potentials to derive relations between the four colligative properties, i.e., (i) relative lowering of vapour pressure, (ii) elevation of boiling point, (iii) depression of freezing point, (iv) Osmotic pressure. Applications in calculating molecular masses of solute; Abnormal colligative properties for associated and dissociated species in solution.
2. Phase rule: Definitions of phase, component and degrees of freedom; Phase rule and its derivation; Phase diagram for water, CO₂, Sulphur
3. First order phase transition and Clapeyron equation; Clausius-Clapeyron equation - derivation and use; Liquid vapour equilibrium for two component systems.

12

Ideal Solution

15

Ideal Solution

16

Ideal & Non-Ideal Solution
Ideal solutions and Non-ideal solutions

Ideal solutions	Non-ideal solutions
The solutions that obey Raoult's Law over the entire range of concentrations are known as ideal solutions.	When a solution does not obey Raoult's Law over the entire range of concentration, then it is called non-ideal solution.
$\Delta_{\text{mix}} H = 0$ and $\Delta_{\text{mix}} V = 0$	$\Delta_{\text{mix}} H \neq 0$ and $\Delta_{\text{mix}} V \neq 0$
The intermolecular attractive forces between solute molecules and solvent molecules are nearly equal to those present between solute and solvent molecules i.e. A-A and B-B interactions are nearly equal to those between A-B.	The intermolecular attractive forces between solute molecules and solvent molecules are not equal to those present between solute and solvent molecules i.e. A-A and B-B interactions are not equal to those between A-B.

17

Vapour Pressure

the pressure exerted by a vapor on a liquid when they are in equilibrium

Vapour Pressure & Boiling Point

At a given temperature, the vapour pressure of a liquid increases with increasing temperature.

Boiling occurs when the vapour pressure of a liquid is equal to the external pressure.

• Normal boiling point
 • Boiling point

20

Vapour Pressure

• Normal boiling point
 • Boiling point

21

Colligative Properties

Colligative properties are characteristics of a solution that depend on the ratio of the number of solute particles to solvent particles.

• Freezing Point Depression
 • Boiling Point Elevation
 • Osmotic Pressure
 • Vapor Pressure Lowering

22

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PDF Sample

Illustrations

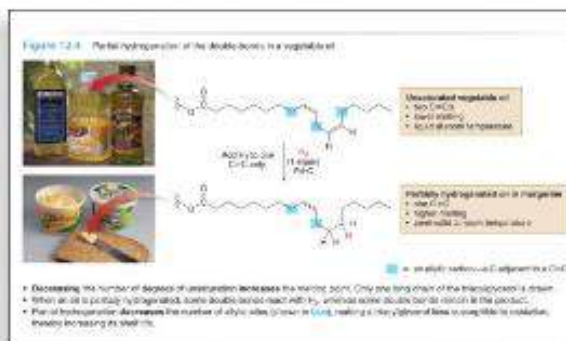
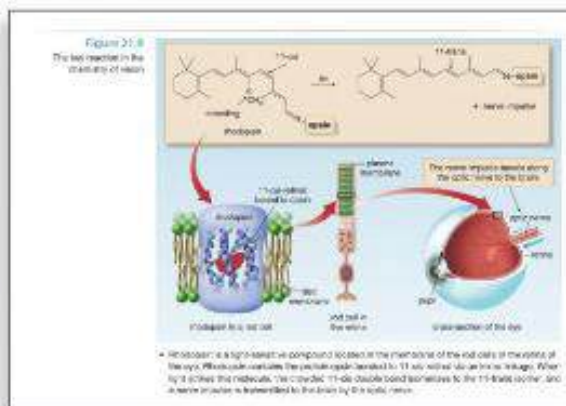
Organic Chemistry is supported by a well-developed illustration program. Besides traditional skeletal (line) structures and condensed formulas, there are numerous ball-and-stick molecular models and electrostatic potential maps to help students grasp the three-dimensional structure of molecules (including stereochemistry) and to better understand the distribution of electronic charge.

"I believe that dissecting the text gives students time to understand and to digest, step-by-step, each concept presented, rather than memorize them. This helps students in achieving better results faster. . . . The quality of the illustrations is very good, without unnecessary explanations that could make them confusing. The language is easy to follow, the concept easy to understand."

—Camelia Gogonea, Cleveland State University

Micro-to-Macro Illustrations

Unique to *Organic Chemistry* are micro-to-macro illustrations, where line art and photos combine with chemical structures to reveal the underlying molecular structures giving rise to macroscopic properties of common phenomena. Examples include starch and cellulose (Chapter 5), adrenaline (Chapter 7), partial hydrogenation of vegetable oil (Chapter 12), and dopamine (Chapter 25).



Class Materials

Slides Outline

- 11 Calculating osmotic pressure
 - The ideal gas law states
 - Assuming $\pi = \frac{P}{RT}$
 - When $\pi = \frac{P}{RT}$
 - Note: osmotic pressure is not a colligative property
- 12 Osmotic pressure and molecular mass
 - Molar mass can be calculated from any of the colligative properties
 - Osmotic pressure provides the most accurate determination because of the magnitude of π (0.001 atm to 1 atm) compared to other colligative properties (e.g., boiling point elevation, freezing point depression, etc.)
- 13 Determining molar mass
 - A solution contains 20.0 mg insulin in 5.00 mL develops an osmotic pressure of 12.5 mm Hg at 300 K
- 14

Determining molar mass

- A solution contains 20.0 mg insulin in 5.00 mL develops an osmotic pressure of 12.5 mm Hg at 300 K

$$M = \frac{\Pi}{RT}$$

$$M = \frac{12.5 \text{ mmHg} \cdot \frac{1}{760 \text{ mmHg}}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 300 \text{ K}} = 6.68 \times 10^{-4} \text{ M}$$

Class taken by Dr. Sucheta Joy, Assistant Professor, department of Chemistry on colligative properties

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34

35

36

37

Molecular Polarizability

The size of the induced dipole μ , depends both on the magnitude of the applied field, E , and on the ease with which the molecule can be distorted.

$$\mu = \alpha E$$

where α is the polarizability of the molecule.

Polarizability is anisotropic

Least polarizability

Greatest polarizability

- Electrons forming the bond are less easily displaced by the field across the bond axis
- Electrons forming the bond are more easily displaced by the field along the bond axis

Class taken by Dr. Debasmita Sardar, Assistant Professor, department of Chemistry on Raman Spectroscopy

The Quantum Mechanical Model of the Atom

THERE IS A SET OF FIVE DIFFERENT *D* ORBITALS.

THERE IS A SET OF SEVEN *F* ORBITALS.

EACH ORBITAL REGARDLESS OF ITS SHAPE HOLDS 2 ELECTRONS.

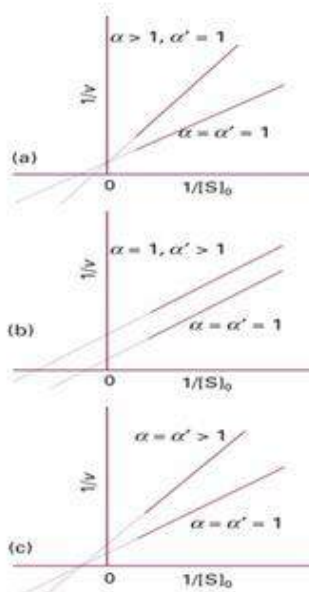
d orbitals

f orbitals

Figure 3-17
Investigating Chemistry, First Edition
© 2007 W. H. Freeman and Company

Class taken by Dr. Dr. Sucheta Joy, Assistant Professor, department of Chemistry on atomic structure

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- Lineweaver–Burk plots characteristic of the three major modes of enzyme inhibition:
- (a) competitive inhibition,
- (b) uncompetitive inhibition, and
- (c) non-competitive inhibition, showing the special case $\alpha = \alpha' > 1$.

$$[E]_0 = \frac{K_M [ES]}{[S]_0} \alpha + [ES] \alpha' = [ES] \left(\frac{\alpha K_M}{[S]_0} + \alpha' \right)$$

Class taken by Dr. Sucheta Joy, Assistant Professor, department of Chemistry on photochemistry


Slides Outline

1. Introduction
2. Green Chemistry Principles
3. Prevention of Waste/By-Products
4. The Role of Green Chemistry
5. Green Chemistry in Industry
6. Green Chemistry in Education
7. Green Chemistry in Society

1. Prevention of Waste/By-Products

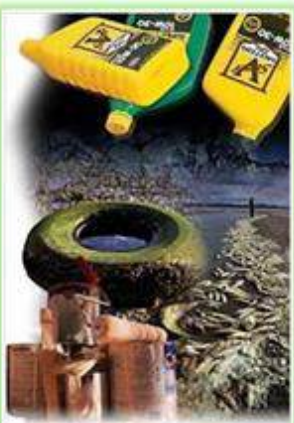
It is better to prevent waste/By Products than to treat or clean up waste after it is formed.

- ✓ Carry out a synthesis in such a way so that formation of waste (by-products) is minimum or absent.
- ✓ Cost involved in the treatment and disposal of waste adds to the overall production cost.
- ✓ The unreacted starting materials also form part of the waste.
- ✓ If discharged causes pollution and requires expenditure for cleaning-up.



❖ An Ounce of Prevention is Worth a Pound Of Cure.

❖ A Chemist in a Green Chemistry lab is Performing Pollution Prevention on a Molecular level .



Class taken by Mrs. Subhra Dholey, SACT, department of Chemistry on green chemistry

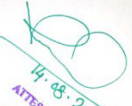
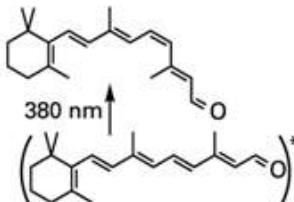

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 Chhatrapati, Tirunelveli

Table 23.1 Examples of photochemical processes

Process	General form	Example
Ionization	$A^* \rightarrow A^+ + e^-$	$\text{NO}^* \xrightarrow{134 \text{ nm}} \text{NO}^+ + e^-$
Electron transfer	$A^* + B \rightarrow A^+ + B^- \text{ or } A^- + B^+$	$[\text{Ru}(\text{bpy})_3]^{2+} + \text{Fe}^{3+} \xrightarrow{452 \text{ nm}} [\text{Ru}(\text{bpy})_3]^{3+} + \text{Fe}^{2+}$
Dissociation	$A^* \rightarrow B + C$	$\text{O}_3^* \xrightarrow{1180 \text{ nm}} \text{O}_2 + \text{O}$
	$A^* + B-C \rightarrow A + B + C$	$\text{Hg}^* + \text{CH}_4 \xrightarrow{254 \text{ nm}} \text{Hg} + \text{CH}_3 + \text{H}$
Addition	$2 A^* \rightarrow B$	$2 \left(\text{CH}_2=\text{CH}-\text{CH}=\text{CH}_2 \right)^* \xrightarrow{230 \text{ nm}} \text{Cyclobutane}$
Abstraction	$A^* + B \rightarrow AB$	$\text{Hg}^* + \text{H}_2 \xrightarrow{254 \text{ nm}} \text{HgH} + \text{H}$
Isomerization or rearrangement	$A^* \rightarrow A'$	

* Excited state.

Table 23-1
Atkins Physical Chemistry, Eighth Edition
 © 2006 Peter Atkins and Julio de Paula

Class taken by Dr. Sucheta Joy, Assistant Professor, department of Chemistry on photochemical process

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Scanned hand written notes

For anharmonic oscillator, find the expression for dissociation energy (D_0) and dissociation energy (D_e) in terms of $\bar{\omega}_e$ and x_e .

Quantised vibrational energy

$$E_v = (v + \frac{1}{2})\bar{\omega}_e - (v + \frac{1}{2})^2 \bar{\omega}_e x_e \quad (1)$$

For max value of quantisation Energy

$$\frac{dE_v}{dv} = 0 \quad (2)$$

From (1) $\frac{dE_v}{dv} = \bar{\omega}_e - 2(v + \frac{1}{2}) \cdot \bar{\omega}_e x_e$

$$= \bar{\omega}_e - 2v \bar{\omega}_e x_e - \bar{\omega}_e x_e$$

$$= \bar{\omega}_e (1 - 2v x_e - x_e) \quad [\bar{\omega}_e \neq 0]$$

From (1) & (2) $1 - 2v x_e - x_e = 0$

$$1 - x_e = 2v x_e \quad \therefore v_{\max} = \frac{1 - x_e}{2x_e} = \frac{1}{2x_e} - \frac{1}{2}$$

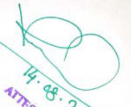
$$E_{v, \max} = (v_{\max} + \frac{1}{2})\bar{\omega}_e - (v_{\max} + \frac{1}{2})^2 \bar{\omega}_e x_e = (\frac{1}{2x_e} - \frac{1}{2} + \frac{1}{2})\bar{\omega}_e - (\frac{1}{2x_e} - \frac{1}{2} + \frac{1}{2})^2 \bar{\omega}_e x_e$$

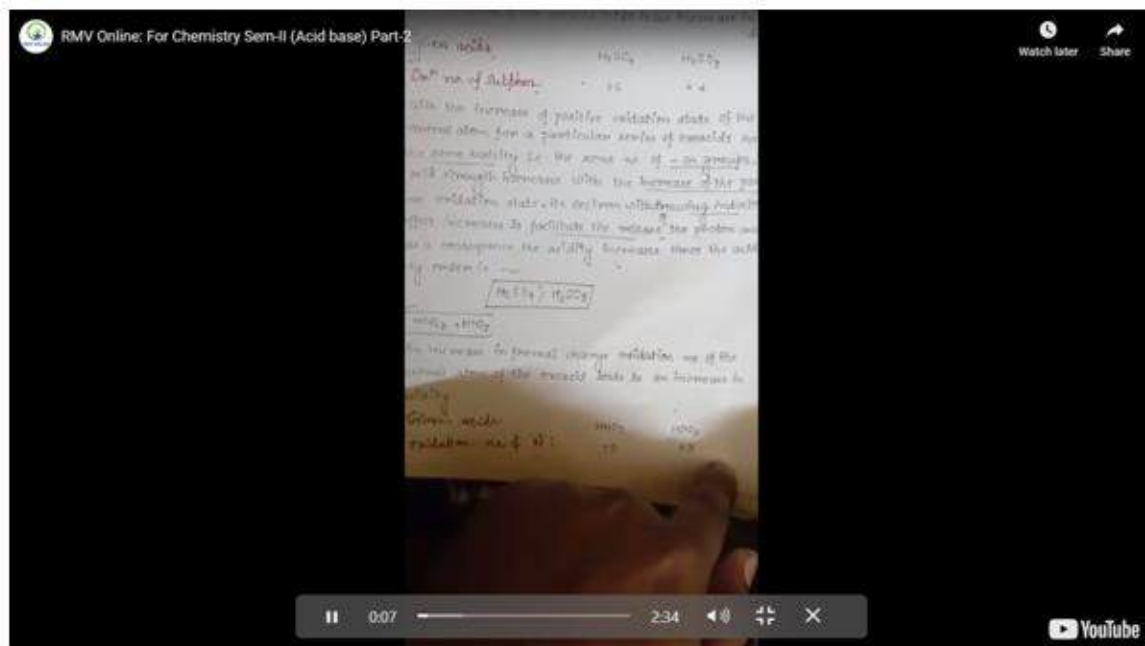
$$E_{v, \max} = \frac{\bar{\omega}_e}{4x_e} = D_e$$

$$D_0 = D_e - E_{\text{zero pt}} = \frac{\bar{\omega}_e}{4x_e} - [\frac{1}{2}\bar{\omega}_e - \frac{1}{4}\bar{\omega}_e x_e]$$

$$D_0 = \frac{\bar{\omega}_e}{2x_e} - \frac{1}{2}\bar{\omega}_e (1 - \frac{1}{2}x_e)$$

Class taken by Dr. Debasmitta Sardar, Assistant Professor, department of Chemistry on
IR Spectroscopy


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 Chittagong, Pabna



Class taken by Dr. Rabiul Alam, Assistant Professor, department of Chemistry on acid-base

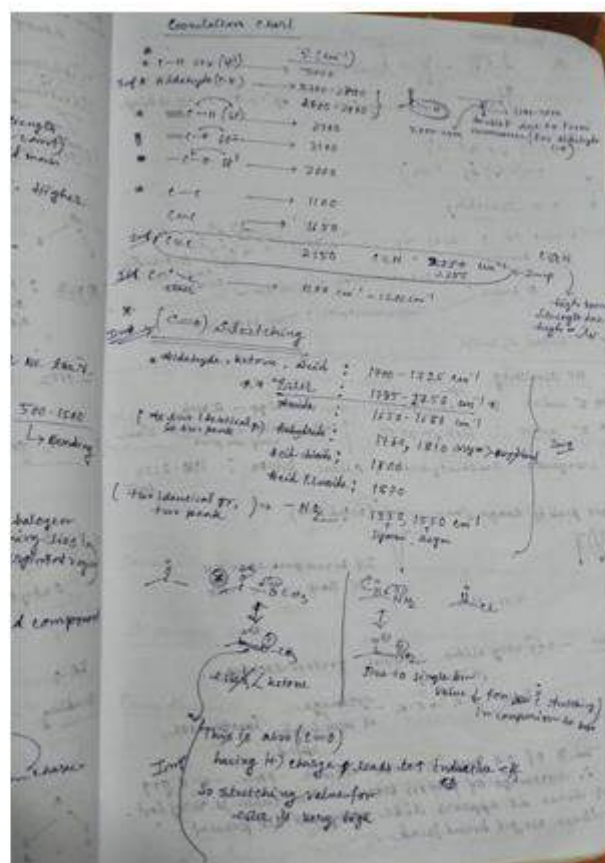
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Rabiul Alam

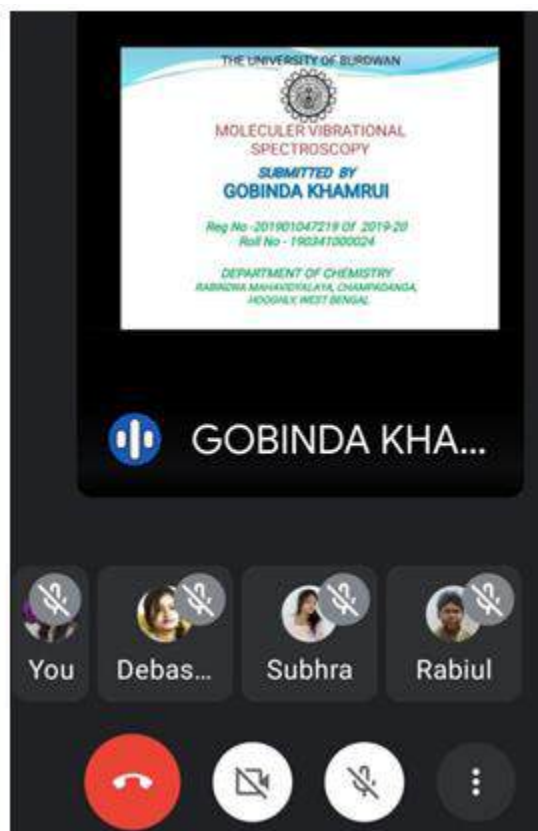
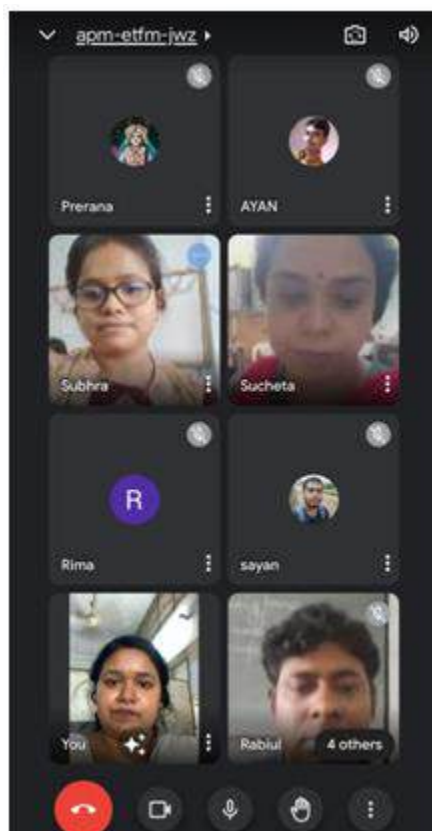
Chandigarh, India



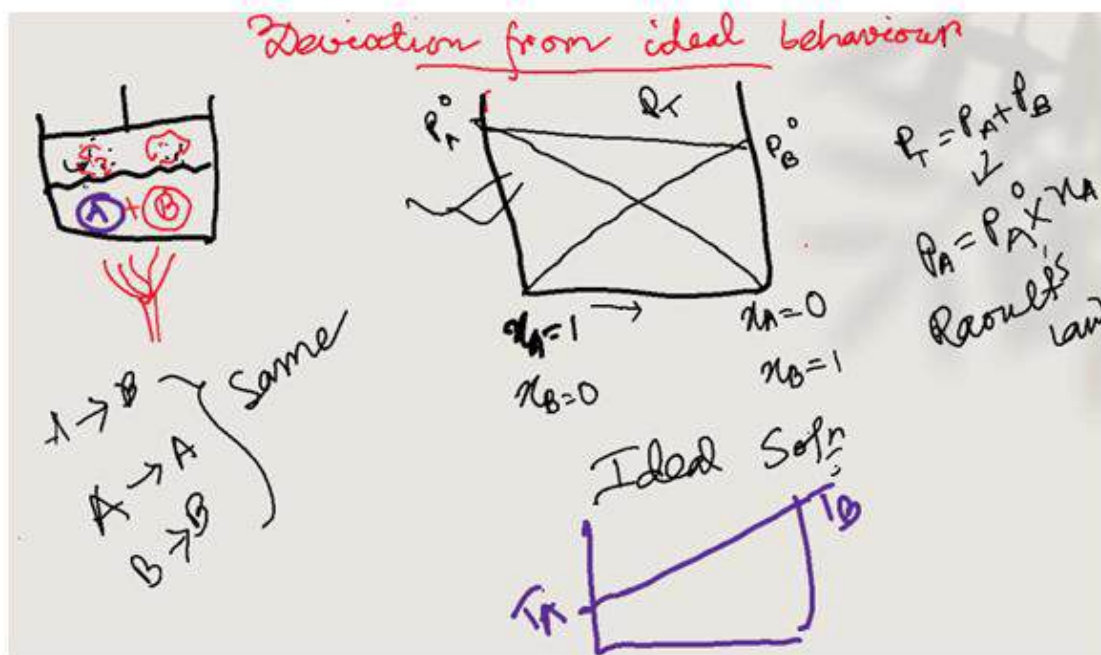
Class taken by Mrs. Subhra dholey, SACT, department of Chemistry on stretching frequency

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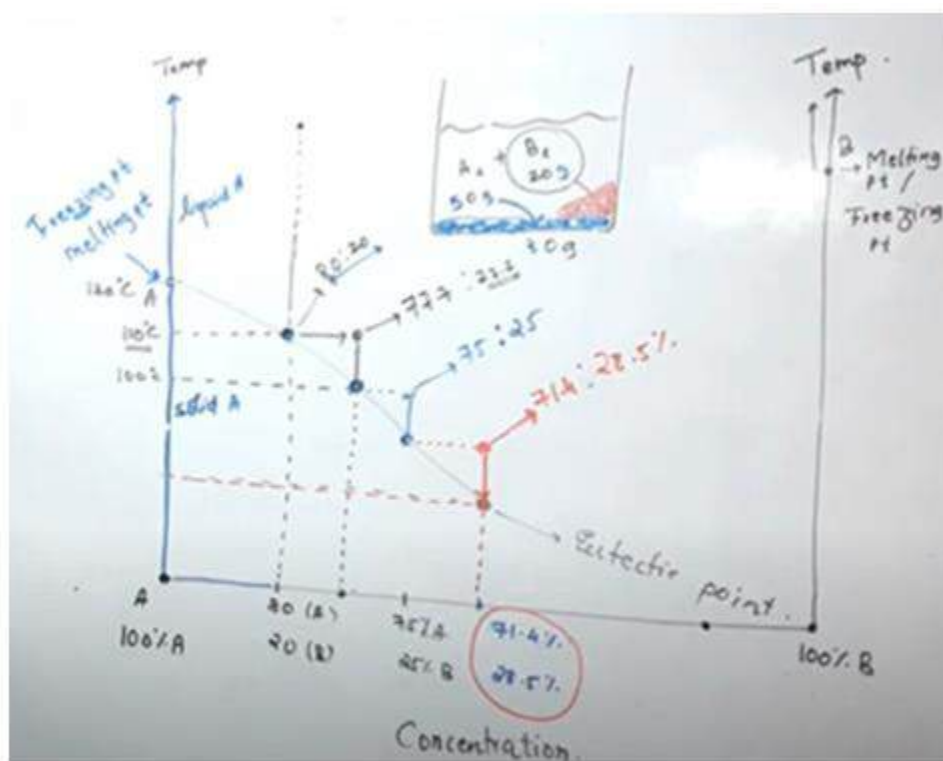
Google Classroom



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Class taken by Dr. Debasmita Sardar, Assistant Professor, department of Chemistry on solutions




Class taken by Dr. Debasmita Sardar, Assistant Professor, department of Chemistry on phase equilibrium

Students Review for online classes & uploaded class notes

 RMV Online

Home Library Login

Department of Chemistry



Teaching Learning Materials for Chemistry

By [Abir Mondal](#) on 2020-05-29 06:34:21

Online class is very very helpful and effective at the vital time

By [Abir Mondal](#) on 2020-05-29 06:33:42

Online class is very very helpful and effective at the vital time

By [Abir Mondal](#) on 2020-05-28 07:51:55

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By [Abir Mondal](#) on 2020-05-28 07:50:59

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By [Monisankar Mukherjee](#) on 2020-05-14 07:20:23

Thank u sir and ma'am .

← SJ class

Name ↑

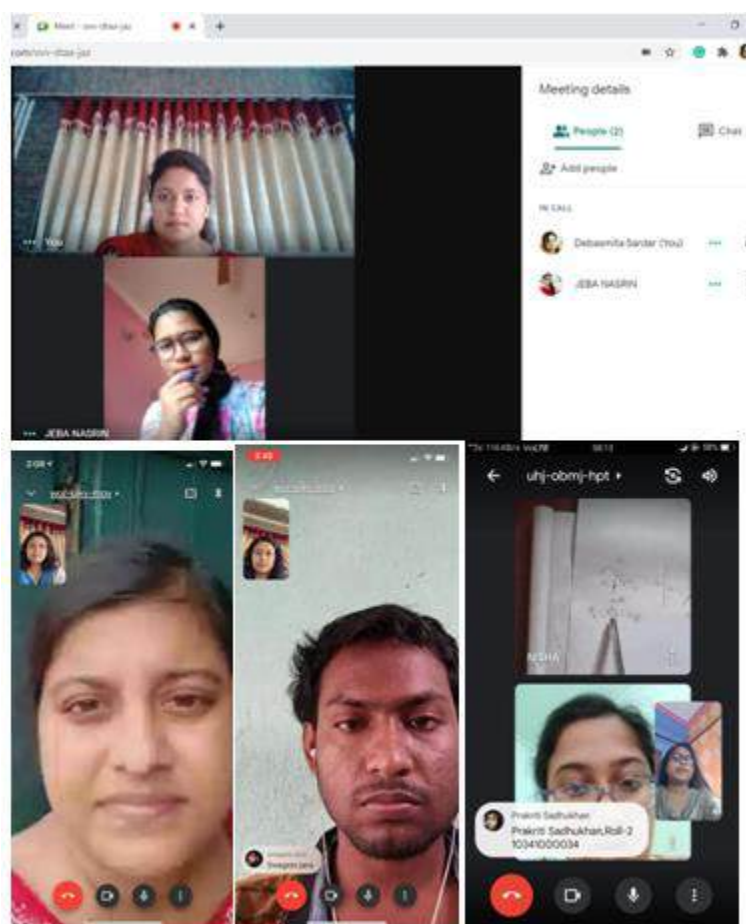
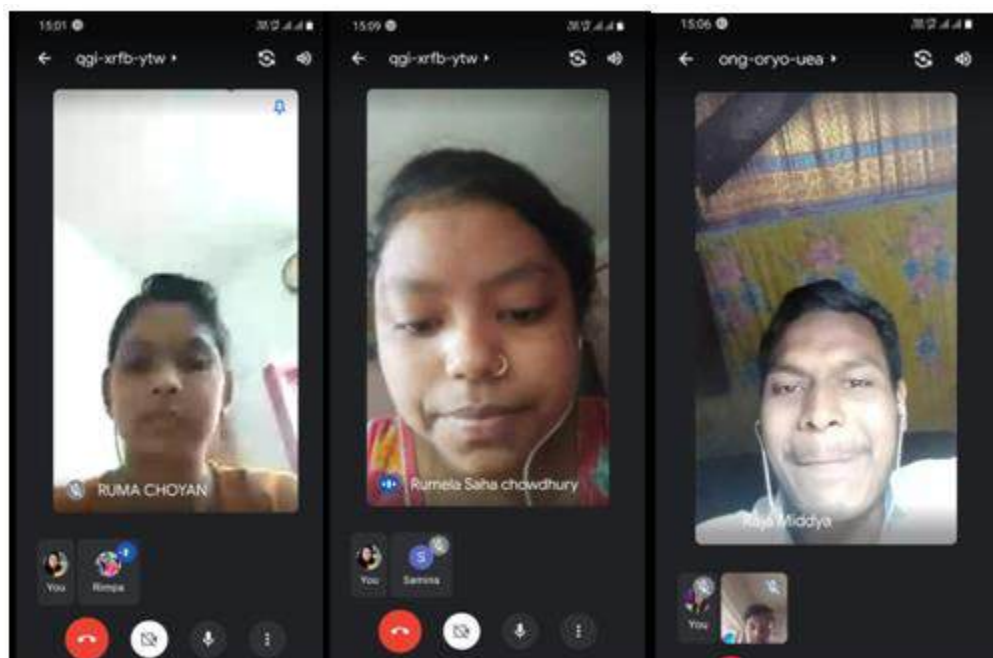
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Modified 30 May 2020
- Absorption-2-02. 06.2020
Modified 9 Jun 2020
- Absorption-3-05. 06.2020
Modified 9 Jun 2020
- Photochemistry-1-05. 05.2020
Modified 7 May 2020
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Modified 7 May 2020
- Photochemistry-3-12. 05.2020
Modified 16 May 2020
- Photochemistry-4-14. 05.2020
Modified 16 May 2020

+

Page Spectroscopy: 1-28. 04.20

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Exam Viva



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Chairman, Tirohity

Practical Exam Viva (in presence of External)

← sfm-yubr-wbp ▶ 🔍 🔊

sayan

Fe₄[Fe(m)₆]
+ 1x(1+1) + 1xX + 6x(1) = 0
4x(1+1) + 1xX - 6 = 0
X = +2

Tanusri

Asita...

Ananya Sadhukhan
Tripti Sadhukhan
Roll 200341000077

Ananya

📞 📺 🎤 ⋮

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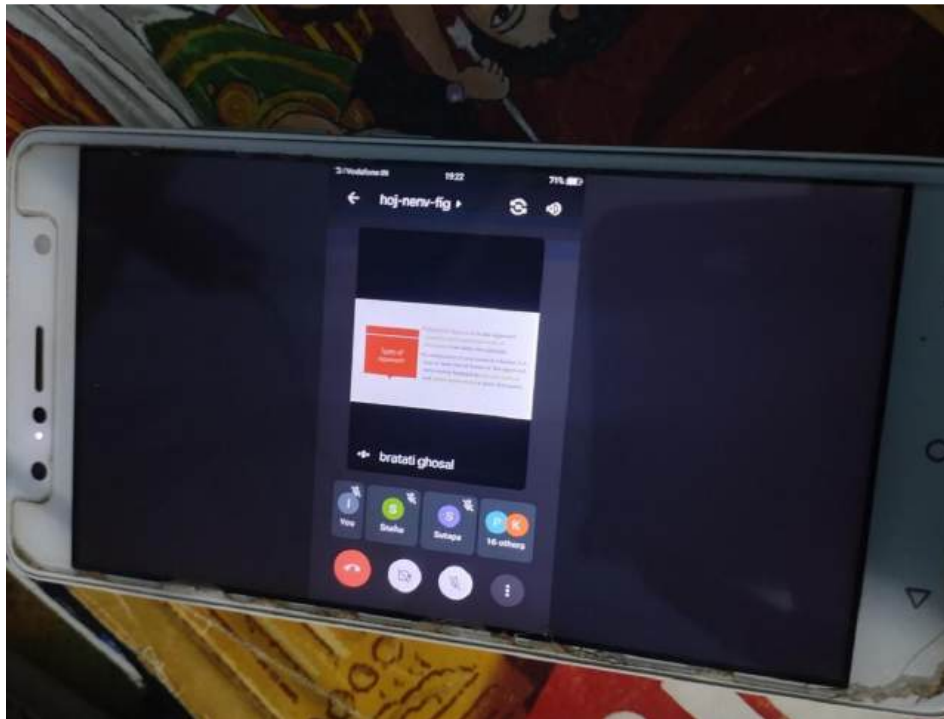
Department of Political Science



Online class taken by Bratati Ghosal, Assistant professor in Political Science for 6 th sem honours students (2020-21 session)

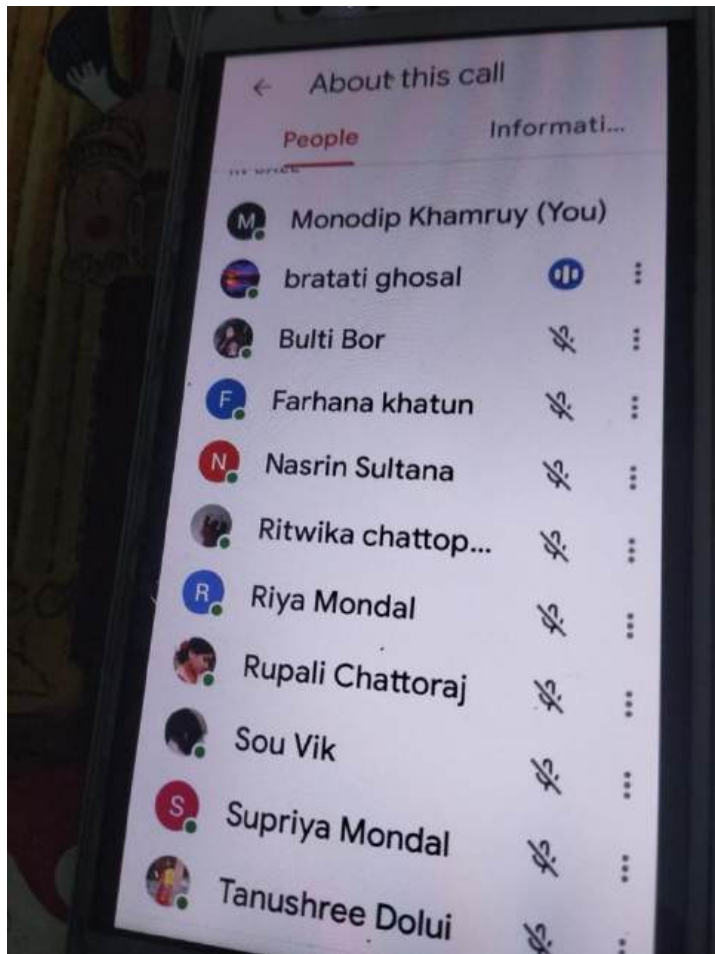
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Chandrasekhar, Tirohly

Online class for Semester-1 Hons students, taken by Bratati Ghosal, Assistant Prof.in Political Sc. (2020-21)



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Online class for Political Science Sem-5 Hons students taken by Bratati Ghosal, Asst.professor for 2020-21



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DEPARTMENT OF PHYSICAL EDUCATION (SESSION 2020-2021)

ONLINE CLASS - SEMESTER-1

CLASS TAKEN BY MR. ARABINDA MAITY , SACT, PHYSICAL EDUCATION



































12:33 12:33 94%

← About this call

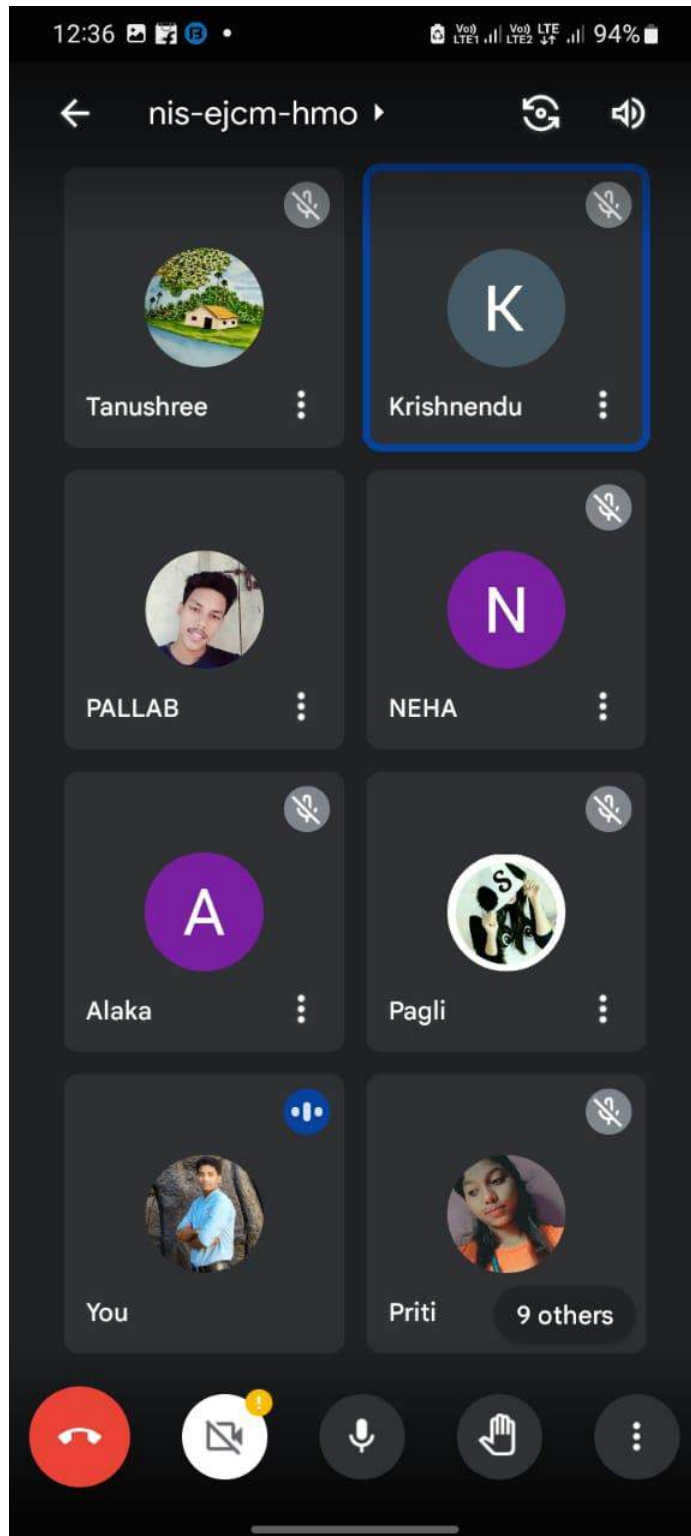
People Information

Search for someone

In call

	Arabinda Maity (You) Meeting host		
	Alaka Hemram		
	Farjana Khatun		
	Krishnendu Mudi		
	NEHA GHARA		
	Pagli P		
	PALLAB MANDAL		
	Priti Poddar		
	Rubina Parvin		
	Sima Adak		
	Smritikana Das		
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14.08.2024
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Topic: Renaissance in Italy.

Semester- III Honours

Paper- CC- VI

Lecture delivered by Prof. *Sujata Bandyopadhyay*, Department of History, Rabindra Mahavidyalaya, Champadanga in 2020.

<https://www.youtube.com/watch?v=EJZaz7YRsiw>

Rests of the departments also have followed the similar techniques to conduct classes and evaluate student performance. (Screenshots are not attached.)



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